

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) for the reaction(s) described. Show the calculation for pH or pOH. Write the chemical formula for the acid or base.

1. What is the pH of a 0.1 M solution of hydrochloric acid?
2. What is the pH of a 0.1 M solution of sodium hydroxide?
3. What is the pH of a 0.1 M solution of acetic acid?
4. What is the pH of a 0.1 M solution of ammonia?
5. What is the pH of a 0.1 M solution of sulfuric acid?
6. What is the pH of a 0.1 M solution of nitric acid?
7. What is the pH of a 0.1 M solution of hydrofluoric acid?
8. What is the pH of a 0.1 M solution of carbonic acid?
9. What is the pH of a 0.1 M solution of phosphoric acid?
10. What is the pH of a 0.1 M solution of hydrocyanic acid?

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8. What is the pH of a 0.1 M solution of carbonic acid?
9. What is the pH of a 0.1 M solution of phosphoric acid?
10. What is the pH of a 0.1 M solution of hydrocyanic acid?

19 2 Hydrogen Ions And Acidity Work Sheet Answers

SJ Ball



19 2 Hydrogen Ions And Acidity Work Sheet Answers:

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