

GENERAL

- Organic molecules : contains C, forms backbone and branched structures of large diverse molecules as rings or chains, **MUST** form 4 covalent bonds
- Hydrocarbons : contains C & H & sometimes O, alkanes (C_nH_{2n+2}), alkenes (C_nH_{2n}), alkynes (C_nH_{2n-2})

1.1 THE CHEMISTRY OF LIFE

- Know what intramolecular and intermolecular forces are

Intermolecular Forces: attraction **BETWEEN** molecules (e.g. hydrogen bonding), like "outer"

Intramolecular Forces: attractions/forces within the molecules (e.g. ionic, polar covalent and covalent bonding)

- Know the difference between covalent and ionic

Ionic: e^- transfer; where resulting ions attract (metal \rightarrow non-metal)

Covalent: Atoms share e^- to fill valence shell for stable octet (generally 2 non-metals)

- Know what electronegativity is and what electronegativity difference is

Electronegativity: The measure of an atom's attraction to shared electrons

EN Difference: Determines what type of bond there is

- 0.0 – 0.4 : slightly/non-polar covalent (electronically symmetrical, X : X)
- 0.4 – 1.7 : polar covalent (partial charges, S^- & S^+)
- 1.7 + : ionic bond (full charges, $M^+ X^-$)

- Know the difference between polar and nonpolar

Polarity - partial pos or neg charge @ ends of molecule (one with most EN has partial neg)

Polar: An unequal sharing of electrons; one atom attracts the electrons more than the other does

Non-polar: balanced out, usually gases

- **Hydrogen bond** – the attractive force between the partially charged H of one molecule & a partially negative charged atom in another molecule
- Water molecule has bent shape due to electrons on O pushing H's down 

- Know what an isotope and radioisotope is

Isotope: # of p^+ same, # of n^0 different, thus the mass # is different

Radioisotope: Radioactive version, decomposes over time

Bio Study Guide Biochemistry

Marcy Osgood, Karen Ocorr

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Table of Contents Bio Study Guide Biochemistry

1. Understanding the eBook Bio Study Guide Biochemistry
 - The Rise of Digital Reading Bio Study Guide Biochemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Bio Study Guide Biochemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Bio Study Guide Biochemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Bio Study Guide Biochemistry
 - Personalized Recommendations
 - Bio Study Guide Biochemistry User Reviews and Ratings
 - Bio Study Guide Biochemistry and Bestseller Lists
5. Accessing Bio Study Guide Biochemistry Free and Paid eBooks
 - Bio Study Guide Biochemistry Public Domain eBooks
 - Bio Study Guide Biochemistry eBook Subscription Services
 - Bio Study Guide Biochemistry Budget-Friendly Options

6. Navigating Bio Study Guide Biochemistry eBook Formats
 - ePUB, PDF, MOBI, and More
 - Bio Study Guide Biochemistry Compatibility with Devices
 - Bio Study Guide Biochemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Bio Study Guide Biochemistry
 - Highlighting and Note-Taking Bio Study Guide Biochemistry
 - Interactive Elements Bio Study Guide Biochemistry
8. Staying Engaged with Bio Study Guide Biochemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Bio Study Guide Biochemistry
9. Balancing eBooks and Physical Books Bio Study Guide Biochemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Bio Study Guide Biochemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Bio Study Guide Biochemistry
 - Setting Reading Goals Bio Study Guide Biochemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Bio Study Guide Biochemistry
 - Fact-Checking eBook Content of Bio Study Guide Biochemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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