

# ACIDS, BASES AND SALTS

<b>Definition of an acid</b> An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	<b>What causes the acidic properties of acids?</b> The hydrogen ions ( $H^+$ ) cause the acidic properties and these are formed in the presence of water.
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<b>Another term to refer to an acid</b> An acid is called a proton donor	<b>Why an acid is also called a proton donor?</b> It's because an acid provides protons or hydrogen ions ( $H^+$ ) to other substances during the reaction.
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<b>Substance to which an acid provides protons</b> BASE	<b>A base is then called</b> PROTON ACCEPTOR	<b>Why the base is called a proton acceptor?</b> Because the base accepts hydrogen ions from acids	<b>Equation for the reaction between acid and base</b> $H_{(aq)} + OH_{(aq)} \longrightarrow H_2O_{(l)}$
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<b>Common laboratory acids</b>		<b>These three common laboratory acids are also called</b> MINERAL ACIDS	<b>Why these acids are also called mineral acids?</b> They are derived from mineral salts ie chlorides for $HCl$ , sulphates for $H_2SO_4$ and nitrates for $HNO_3$
Hydrochloric acid	$HCl$		
Sulphuric acid	$H_2SO_4$		
Nitric acid	$HNO_3$		

<b>Other mineral acids known</b>	<b>Mineral salts from which the acid is derived</b>	<b>Organic acids known</b>	<b>Naturally occurring acids</b>
Sulphurous acid	$H_2SO_3$	Derived from sulphites	CITRIC ACID from lemons
Carbonic acid	$H_2CO_3$	Derived from carbonates	TARTARIC ACID from grapes
Phosphoric acid	$H_3PO_4$	Derived from phosphates	ACETIC ACID from vinegar
Nitrous acid	$HNO_2$	Derived from nitrates	LACTIC ACID from sour milk
		Methanoic acid ( $HCOOH$ )	Hydrochloric acid from digestive juices

<b>Whenever an acid is dissolved in water, it produces</b> HYDROGEN IONS	<b>Term given to the number of hydrogen ions produced by one molecule of an acid</b> BASICITY OF AN ACID	<b>Definition of basicity of an acid</b> BASICITY of an acid is the <u>number of hydrogen ions produced by one molecule of an acid in aqueous solution</u> .
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# Acids Bases And Salts Study Guide Awnsers

**M Tight**

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