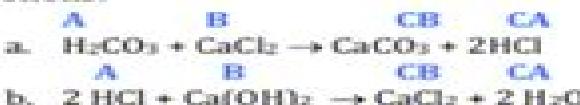


## Chapter 19 Study Guide: Acids, Bases & Titration – Answers

1. Differentiate between Arrhenius, Brønsted-Lowry & Lewis acids and bases.

	<b>ACIDS</b>	<b>BASES</b>
<b>Arrhenius</b>	Form hydronium ions ( $\text{H}_3\text{O}^+$ ) in aqueous sol'n	Form hydroxide ions ( $\text{OH}^-$ ) in aqueous sol'n
<b>Brønsted-Lowry</b>	Proton ( $\text{H}^+$ ) donors	Proton ( $\text{H}^+$ ) acceptors
<b>Lewis</b>	Electron pair acceptors	Electron pair donors

2. Label the acid (A), base (B), conjugate acid (CA) and conjugate base (CB) in each of the following reactions:



3. Write the names or formulas for the following acids and bases:

a. $\text{H}_2\text{S}$ – <b>hydrosulfuric acid</b>	e. Nitrous acid – <b><math>\text{HNO}_2</math></b>
b. $\text{HC}_2\text{H}_3\text{O}_2$ – <b>acetic acid</b>	f. Calcium hydroxide – <b><math>\text{Ca}(\text{OH})_2</math></b>
c. $\text{NaOH}$ – <b>sodium hydroxide</b>	g. hydrochloric acid – <b><math>\text{HCl}</math></b>
d. $\text{HNO}_3$ – <b>nitric acid</b>	

Use the following formulas for the remaining questions: (these will not be given to you on the test!)

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$$

$$\text{pH} = -\log[\text{H}_3\text{O}^+]$$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$\text{pH} + \text{pOH} = 14$$

$$\text{M}_\text{A}\text{-V}_\text{A} \cdot \text{n}_\text{A} = \text{M}_\text{B}\text{-V}_\text{B} \cdot \text{n}_\text{B}$$

4. What is the pH of  $1.34 \times 10^{-4}$  M solution of hydrochloric acid?

$$\text{pH} = -\log[1.34 \times 10^{-4}\text{M}] = 3.873$$

5. Determine the pH and pOH of a  $1.85 \times 10^{-6}$  M NaOH solution.

$$\text{pOH} = -\log[1.85 \times 10^{-6}\text{M}] = 5.733$$

$$\text{pH} = 14 - 5.73 = 8.267$$

6. What is the  $[\text{OH}^-]$  of an aqueous solution that has a pH of 5.0?

$$\text{pOH} = 9 \quad [\text{OH}^-] = 1 \times 10^{-9}\text{M}$$

7. What is the pH and the pOH of a solution with a volume of 2000. mL containing 35 grams of NaOH? Is this solution acidic or basic?

$$\frac{35\text{gNaOH}}{2.000\text{L}} \times \frac{1\text{molNaOH}}{40.00\text{gNaOH}} = 0.44\text{M NaOH}$$

$$\text{pOH} = -\log[0.44\text{M}] = 0.36 \quad \text{pH} = 14 - 0.36 = 13.64 \quad \text{basic}$$

# Acids And Bases Final Study Guide Answers

**John Michels (Journalist)**

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in the following dosage forms tablet solution

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