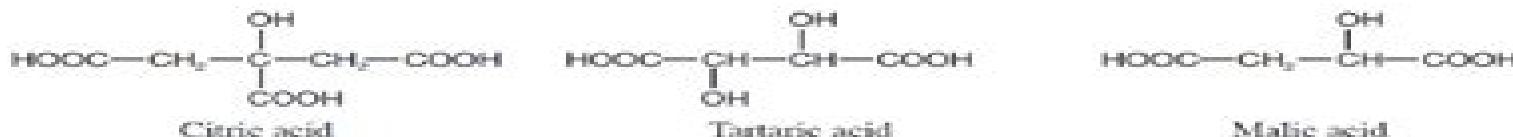


## Acidity of Beverages Lab

**Introduction:** Common beverages may be either acidic or basic. Fruit juices, for example, get their sweet taste from sugars and their sour or tart taste from weak acids such as citric acid. If the juice contains too much sugar, it will taste bland, but too much acid and the juice will taste sour. The concentration of acids in various consumer beverages may be determined by titration with sodium hydroxide.

**Background:** The main acids present in fruits and fruit juices are citric acid (in citrus fruits), tartaric acid (in grapes), and malic acid (in apples). All of these are characterized as weak acids.



**Figure 1.** Organic acids in fruits and fruit juices

The amount of citric acid in citrus fruit juices can be determined by titration with a standard solution of sodium hydroxide. A standard solution is one whose concentration is accurately known, usually to three significant figures. Citric acid is a tricarboxylic acid—it has three ionizable or “active” hydrogen atoms in its structure. One mole of citric acid therefore reacts with three moles of sodium hydroxide via the acid-base neutralization reaction shown in Equation 1.



Acid-base titrations are an extremely useful technique to determine the concentration of an acid or base in a sample. In titrating beverages such as orange juice, apple juice, and sodas that contain weak acids, the juice is called the analyte and a strong base is used as the titrant.

In the titration procedure, a sodium hydroxide solution of known molarity is carefully added using a buret to a measured volume of fruit juice containing an indicator. The exact volume of sodium hydroxide that must be added to reach the indicator endpoint is measured and then used to calculate the concentration of citric acid in the juice.

A sample setup for a titration is shown in Figure 2 on the next page, where a buret containing the titrant is clamped to the support stand and a beaker or flask containing the analyte is set on a stir plate. If a pH probe is inserted into the solution, a titration curve can be constructed by plotting the pH of the solution on the y-axis versus the volume of titrant added on the x-axis. The shape of the titration curve may be used to distinguish strong and weak acids in the analyte, and also permits graphical analysis of the equivalence point. At the equivalence point, moles of added titrant are stoichiometrically related to moles of analyte in the sample.

# Acidity Of Beverages Chem Fax Lab Answers

**Kathleen Armour**

## Acidity Of Beverages Chem Fax Lab Answers:

**Worldwide Biotech Industry Reporting Service** ,2000 Alphabetical listing of biotechnology corporations Entries cover such information as management facilities investments research and development products in development and subject terms Subject geographic and investor indexes    **California Farmer** ,2007    **Consumer Sourcebook** ,2001    Ward's Business Directory of U.S. Private and Public Companies ,2005    Directory of Testing Laboratories ,1994    Who's who in European Research and Development ,1997    Directory of Graduate Research ,2001 Faculties publications and doctoral theses in departments or divisions of chemistry chemical engineering biochemistry and pharmaceutical and or medicinal chemistry at universities in the United States and Canada    Chemical Engineering Progress ,2002    **Food Australia** ,1995    **Companies and Their Brands** ,1990    Consultants & Consulting Organizations Directory Cengage Gale,2009-05-12    *Book of Abstracts* ,2000    **What's Wrong with My Drink** R. H. Molony,2005-01-01

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## **Table of Contents Acidity Of Beverages Chem Fax Lab Answers**

1. Understanding the eBook Acidity Of Beverages Chem Fax Lab Answers
  - The Rise of Digital Reading Acidity Of Beverages Chem Fax Lab Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Acidity Of Beverages Chem Fax Lab Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acidity Of Beverages Chem Fax Lab Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Acidity Of Beverages Chem Fax Lab Answers
  - Personalized Recommendations
  - Acidity Of Beverages Chem Fax Lab Answers User Reviews and Ratings
  - Acidity Of Beverages Chem Fax Lab Answers and Bestseller Lists

5. Accessing Acidity Of Beverages Chem Fax Lab Answers Free and Paid eBooks
  - Acidity Of Beverages Chem Fax Lab Answers Public Domain eBooks
  - Acidity Of Beverages Chem Fax Lab Answers eBook Subscription Services
  - Acidity Of Beverages Chem Fax Lab Answers Budget-Friendly Options
6. Navigating Acidity Of Beverages Chem Fax Lab Answers eBook Formats
  - ePUB, PDF, MOBI, and More
  - Acidity Of Beverages Chem Fax Lab Answers Compatibility with Devices
  - Acidity Of Beverages Chem Fax Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Acidity Of Beverages Chem Fax Lab Answers
  - Highlighting and Note-Taking Acidity Of Beverages Chem Fax Lab Answers
  - Interactive Elements Acidity Of Beverages Chem Fax Lab Answers
8. Staying Engaged with Acidity Of Beverages Chem Fax Lab Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Acidity Of Beverages Chem Fax Lab Answers
9. Balancing eBooks and Physical Books Acidity Of Beverages Chem Fax Lab Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Acidity Of Beverages Chem Fax Lab Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Acidity Of Beverages Chem Fax Lab Answers
  - Setting Reading Goals Acidity Of Beverages Chem Fax Lab Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acidity Of Beverages Chem Fax Lab Answers
  - Fact-Checking eBook Content of Acidity Of Beverages Chem Fax Lab Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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