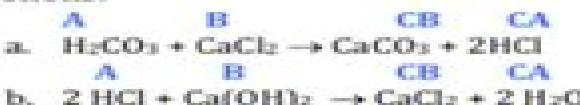


Chapter 19 Study Guide: Acids, Bases & Titration – Answers

1. Differentiate between Arrhenius, Brønsted-Lowry & Lewis acids and bases.

| | ACIDS | BASES |
|-----------------------|---|--|
| Arrhenius | Form hydronium ions (H_3O^+) in aqueous sol'n | Form hydroxide ions (OH^-) in aqueous sol'n |
| Brønsted-Lowry | Proton (H^+) donors | Proton (H^+) acceptors |
| Lewis | Electron pair acceptors | Electron pair donors |

2. Label the acid (A), base (B), conjugate acid (CA) and conjugate base (CB) in each of the following reactions:



3. Write the names or formulas for the following acids and bases:

| | |
|---|---|
| a. H_2S – hydrosulfuric acid | e. Nitrous acid – HNO_2 |
| b. $\text{HC}_2\text{H}_3\text{O}_2$ – acetic acid | f. Calcium hydroxide – $\text{Ca}(\text{OH})_2$ |
| c. NaOH – sodium hydroxide | g. hydrochloric acid – HCl |
| d. HNO_3 – nitric acid | |

Use the following formulas for the remaining questions: (these will not be given to you on the test!)

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$$

$$\text{pH} = -\log[\text{H}_3\text{O}^+]$$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$\text{pH} + \text{pOH} = 14$$

$$\text{M}_\text{A}\text{-Va}^- \text{M}_\text{B} = \text{M}_\text{B}\text{-V}_\text{a}^- \text{M}_\text{A}$$

4. What is the pH of 1.34×10^{-4} M solution of hydrochloric acid?

$$\text{pH} = -\log[1.34 \times 10^{-4}\text{M}] = 3.873$$

5. Determine the pH and pOH of a 1.85×10^{-6} M NaOH solution.

$$\text{pOH} = -\log[1.85 \times 10^{-6}\text{M}] = 5.733$$

$$\text{pH} = 14 - 5.73 = 8.267$$

6. What is the $[\text{OH}^-]$ of an aqueous solution that has a pH of 5.0?

$$\text{pOH} = 9 \quad [\text{OH}^-] = 1 \times 10^{-9}\text{M}$$

7. What is the pH and the pOH of a solution with a volume of 2000. mL containing 35 grams of NaOH? Is this solution acidic or basic?

$$\frac{35\text{gNaOH}}{2.000\text{L}} \times \frac{1\text{molNaOH}}{40.00\text{gNaOH}} = 0.44\text{M NaOH}$$

$$\text{pOH} = -\log[0.44\text{M}] = 0.36 \quad \text{pH} = 14 - 0.36 = 13.64 \quad \text{basic}$$

Acid And Bases Chemistry Study Guide Answers

T. R. Dickson

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Student Study Guide to accompany Chemistry Martin Silberberg,Libby Weberg,2005-01-06 **The Complete Chemistry Study Guide and Note Cards and MCAT** Konstantinos Papadopoulos,2012-07-06 **Organic Chemistry, Study Guide and Solutions Manual** T. W. Graham Solomons,Craig B. Fryhle,2000 On the cover of this book is a Pacific yew tree found in the ancient forests of the Pacific Northwest The bark of the Pacific yew tree produces Taxol found to be a highly effective drug against ovarian and breast cancer Taxol blocks mitosis during eukaryotic cell division The supply of Taxol from the Pacific yew tree is vanishingly small however A single 100 year old tree provides only about one dose of the drug roughly 300 mg For this reason as well as the spectacular molecular architecture of Taxol synthetic organic chemists fiercely undertook efforts to synthesize it Five total syntheses of Taxol have thus far been reported Now a combination of isolation of a related metabolite from European yew needles and synthesis of Taxol from that intermediate supply the clinical demand This case clearly demonstrates the importance of synthesis and the use of organic chemistry It s just one of the many examples used in the text that will spark the interest of students and get them involved in the study of organic chemistry

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Dickson,1999-07-07 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and chemical vision and refining problem solving skills Many chapters in this edition has been rewritten and rearranged to vitalize the topics and to include interesting examples analogies and images

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 - The Rise of Digital Reading Acid And Bases Chemistry Study Guide Answers
 - Advantages of eBooks Over Traditional Books
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 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid And Bases Chemistry Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid And Bases Chemistry Study Guide Answers
 - Personalized Recommendations
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5. Accessing Acid And Bases Chemistry Study Guide Answers Free and Paid eBooks
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 - Interactive Elements Acid And Bases Chemistry Study Guide Answers
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid And Bases Chemistry Study Guide Answers
9. Balancing eBooks and Physical Books Acid And Bases Chemistry Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid And Bases Chemistry Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Acid And Bases Chemistry Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acid And Bases Chemistry Study Guide Answers
 - Fact-Checking eBook Content of Acid And Bases Chemistry Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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