

Acids and Bases

Overview of Acids and Bases

Indicators: Acids and Bases

Acids: Donates its proton to base (proton donor)

Base: Donates its lone pair to form a bond (LBP)

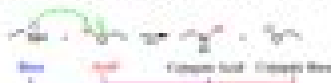


Remember: (1) only H atoms, (2) only OH is a base

What about organic indicators and how organic indicators?

Acids and Bases

Acids are proton donors. (The tendency to bring in more electrons to form a bond.)
Bases are proton acceptors. (The tendency to take away the lone pair.)



General acid-base reactions of acid-base reactions:

to form the more stable, the more polarizable charge of the base

and usually for more polarizable hydrogens of the acid.

to form the more stable of the hydrogens and more the electrons

in the other atom

Remember: **acid-base reactions are reversible**

1. Acid: Donates and Proton: Donates. (or from Proton: Protonates)

What if there is no proton transfer?

Acids and Bases

Acids: acid is an electron-pair acceptor.

Acids: acid is an electron-pair acceptor.



Acid-base reactions: acids and bases are described in terms of electron.

1. One of the acid-base reactions: transfer of a proton

As a general reaction: transfer of a proton from an acid to a base

Acid Strength

What does it mean "strong acid"?

Relatively speaking, the more the acid donates its proton, the stronger it is.



The dissociation of the acid is a reversible reaction that is described by equilibrium constant.



Then: $K_a = \frac{[\text{H}^+][\text{Cl}^-]}{[\text{HCl}]}$

Acid Strength and pK_a

The K_a value of acid dissociation (1) and acid-base (2) are shown below.



The acid is a stronger acid than the base.

To compare the acid, pK_a is used, which is the pK_a of the given acid.

$pK_a = -\log K_a$

$$pK_a = -\log K_a = 10 \quad pK_a = -\log K_a = 14$$

So, the stronger the acid, the lower the pK_a value. (It is to separate the two)

Acid and pK_a values of some common functional groups



What makes the Acid Strong?

The strength of the acid is determined by the stability of the conjugate base. The more stable the conjugate base, the stronger the acid.

The stronger the acid, the more stable the conjugate base.



Acid is a strong acid and its conjugate base is a strong base.

There are two factors to be considered when assessing the stability of the conjugate base.

1. The first and the last factor is the **charge** to which the hydrogens is connected. Usually, an **anion** (or **phenoxide**) and **thiophenoxide**.

2. When the acid and its conjugate base are both charged, usually, an **anion** (or **phenoxide**) and **thiophenoxide**.



Stronger is more electronegative, that is why the acid is more acidic.

$$pK_a = 10 \quad pK_a = 14$$

1. When comparing different acids, the more stable the conjugate base, the stronger the acid.

Large atoms have higher polarizability and usually for negative charge better in the charge is spread out.



There are acid-base reactions that are more electronegative than acids.

The acid is more stable and usually for negative charge better.

2. When comparing for same atom, check which conjugate base is **more electronegative**.



1. The first factor for determining the acidity is called **inductive**.

Inductive is inductive effect is the transmission of charge through atoms.

Electronegativity atoms pull the negative charge and help to stabilize the conjugate base.



The more electronegative, the more electron withdrawing, the stronger the acid.



The closer the electronegative atom to the hydrogens, the stronger the acid.

Increasing Acidity

2. The second factor is the **resonance** of the acid connected to the hydrogens.

The explains the great difference of pK_a values of alcohols, alkenes, and alkenes.



3. Which is more electronegative than O , or the atom S connected, the stronger the acid.

Therefore, for **acids**, **alcohols**, **aldehydes**, **ketones**.

The primary acids: **alcohols**, **aldehydes**, **ketones**, **aldehydes**, **ketones**.

of a conjugate base is **acids**.

Remember: **acid-base reactions are reversible**



The Position of Equilibrium for an Acid-Base Reaction

Is the following reaction reaction possible?



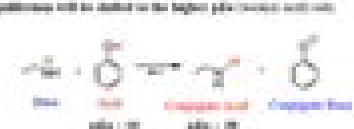
Acid-base reactions are reversible, so neither way of doing the reaction is.

Is the equilibrium of the reaction shifted to the right (products) or to the left (reactants)?

To answer this question, you need to remember: for the equilibrium of acids and bases reaction is shifted towards the weaker acid and base.

So, you need to compare the pK_a of the acid and base and the conjugate acid and base.

The equilibrium will be shifted to the higher pK_a reaction with only.



Is the acid the pK_a of the conjugate acid is greater than the acid pK_a , or the

equilibrium is shifted to the right. Therefore the right pointing arrow is shown larger.

Choosing a Base to Deprotonate a Given Compound

Suppose you need to deprotonate the following compound: **alcohol**

Deprotonate means remove the proton acidic proton.

First, find the pK_a value: **alcohol** $pK_a = 17$

Find the chemical equation of the conjugate reaction with a conjugate base (1).



The pK_a of the conjugate acid must be higher than 17, so weaker acid is formed.

Find a conjugate base with $pK_a > 17$. Acids are weak when $pK_a > 14$.



Choosing an Acid to Protonate a Given Compound

Suppose you need to protonate the following compound: **alcohol** (1) (2) (3) (4) (5) (6) (7) (8) (9) (10) (11) (12) (13) (14) (15) (16) (17) (18) (19) (20) (21) (22) (23) (24) (25) (26) (27) (28) (29) (30) (31) (32) (33) (34) (35) (36) (37) (38) (39) (40) (41) (42) (43) (44) (45) (46) (47) (48) (49) (50) (51) (52) (53) (54) (55) (56) (57) (58) (59) (60) (61) (62) (63) (64) (65) (66) (67) (68) (69) (70) (71) (72) (73) (74) (75) (76) (77) (78) (79) (80) (81) (82) (83) (84) (85) (86) (87) (88) (89) (90) (91) (92) (93) (94) (95) (96) (97) (98) (99) (100)

To do that, write the chemical equation of acid + base pK_a .



The pK_a of the conjugate acid of alcohol is 17. Therefore, the pK_a of **acid**

must be weaker than 17, so stronger acid is formed, weaker and stronger.

Acid is a strong acid and a proton donor (acid) for protonating the acid, a weaker acid

is $pK_a < 17$ and is a stronger acid of an acid than before.



Hammett-Hansch Equation

The equation shows whether the compound with a given pK_a is an acid or base (deprotonated) there is an equilibrium depending on the pK_a of the

$$pK_a = pK_a + \log \frac{[\text{acid}]}{[\text{base}]}$$

One general and application of the Hammett-Hansch equation

1. If the $pK_a < pK_a$ 2. If the $pK_a > pK_a$ 3. If the $pK_a = pK_a$

1. $\log \frac{[\text{acid}]}{[\text{base}]} < 0$ 2. $\log \frac{[\text{acid}]}{[\text{base}]} > 0$ 3. $\log \frac{[\text{acid}]}{[\text{base}]} = 0$

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Acid Base Reaction Study Guide

Jill Meryl Levy



Acid Base Reaction Study Guide:

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