

Acids and Bases

Overview of Acids and Bases

Acids and Bases

Acids dissociate to release **proton** (H^+)

Bases dissociate to accept **proton** (H^+)



Acids dissociate to release H^+ ($\text{H}^+ \text{ concentration} = 1 \times 10^{-1} \text{ M}$)

Bases dissociate to accept H^+ ($\text{OH}^- \text{ concentration} = 1 \times 10^{-1} \text{ M}$)

What does H^+ concentration and OH^- concentration?

Acids and Bases

Acids are proton donors. $\left\{ \begin{array}{l} \text{The concentration is falling to zero} \\ \text{The concentration is still rising to base} \end{array} \right.$



What does concentration of acid/base reactions?

1) At the same time, the base NaOH increases the change of the base and reacts to form acidic hydroxides of the acid.

2) At the end of the reaction, the base decreases.

In the other time

Acid-base reaction occurs when H^+ concentration of acid increases.

2) Both H^+ concentration and OH^- concentration are from **Potential Properties**.

What if there is no proton available?

Acids, Bases, and Salts

Acids react with **alkalis** (bases) to form **salts**.



For the acid-base reaction, acids and bases are described in terms of **strength**.

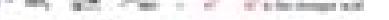
1) One of the acid-base reaction, **stronger** to **weaker**

For a potential reaction, **stronger** acid is H^+ and **weaker** is OH^- .

Acid Strength

What does a strong "strong acid"?

Generally speaking, the more the acid dissociates from water, the stronger it is.



Acid Base Reaction Study Guide

Jill Meryl Levy

Acid Base Reaction Study Guide:

Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder's tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

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