

**Chemical Reactions:**

1. The following reaction  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  is given. What is the  $\Delta H^\circ_f$  for the reaction?

A) greater than 2.00 kcal/mol  
 B) greater than 1.00 kcal/mol  
 C) less than 1.00 kcal/mol  
 D) less than 2.00 kcal/mol  
 E) greater than 4.00 kcal/mol

2. A 1.0-mol sample of  $\text{PCl}_3$  reacts with excess oxygen according to the equation:  $2\text{PCl}_3(\text{g}) \rightarrow 2\text{PCl}_5(\text{g}) + \text{O}_2(\text{g})$ . Below, which reaction will yield the greatest amount of  $\text{O}_2$  after  $\text{PCl}_3$  is removed?

A) 0.0 mol  
 B) 1.0 mol  
 C) 1.2 mol  
 D) 1.7 mol  
 E) 2.0 mol

3. The reaction of  $\text{NO}_2$  is step characterized by the equation:  $4\text{NO}_2 \rightarrow 3\text{O}_2 + 2\text{N}_2\text{O}_5$ . If 1.0 mol of  $\text{NO}_2$  is added to 1.0 mol of  $\text{N}_2\text{O}_5$  and the reaction is allowed to proceed, how will the reaction rate change?

A) 1.000 mol/l  
 B) 0.750 mol/l  
 C) 1.000 mol/l  
 D) 1.50 mol/l  
 E) 2.00 mol/l

4. Calculated are the following reactions, or from the experiments, are unknowns, calculated.

$\text{H}_2\text{O}_2(\text{l}) + \text{H}_2\text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O}_2(\text{l}) = 300\text{J/g}$

$\text{H}_2 + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$  (enthalpy of difference,  $\Delta H^\circ_f$  for production of 1.00 mol of  $\text{H}_2\text{O}_2$ , which is the difference of 200)

A)  $\text{O}_2$   
 B)  $\text{H}_2$   
 C)  $\text{H}_2$   
 D)  $\text{O}_2$   
 E) C

5. Under the following conditions,  $\text{NO}_2(\text{g}) + \text{NO}_2(\text{g}) \rightarrow 2\text{NO}(\text{g}) + \text{O}_2(\text{g})$  is

higher energy molecules of  $\text{NO}_2$  are produced in?

A) 1.000 mol/l  
 B)  $1.000 \times 10^3$   
 C)  $1.000 \times 10^6$   
 D)  $1.000$   
 E) same as above

6. The following reaction  $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$  is performed. Below, which number of molecules would be produced from 1.000 mol of hydrogen and excess oxygen?

A) 1.000 mol  
 B) 2.00 mol  
 C) 0.500 mol  
 D) 0.100 mol  
 E) 0.000 mol

7. Under the following conditions,  $\text{CH}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}(\text{g}) + \text{H}_2\text{O}(\text{l})$ . What mass of oxygen,  $\text{O}_2(\text{g})$  is required to react completely with  $1.0 \times 10^{-3}$  mol  $\text{CH}_4(\text{g})$ ?

A)  $3.00 \times 10^{-3}$  g  
 B)  $3.00 \times 10^{-2}$  g  
 C)  $4.00 \times 10^{-3}$  g  
 D)  $4.00 \times 10^{-2}$  g  
 E)  $0.100 \times 10^{-3}$  g

8. Under reactor conditions and quantities of reactants, the following reaction of  $\text{CH}_4$  with propane is carried out. How many molecules of  $\text{CH}_4$  are required to react completely with  $1.0 \times 10^{-3}$  mol  $\text{CH}_3\text{CH}_2\text{CH}_3$ ?



A)  $1.000 \times 10^{10}$  molecules  
 B)  $1.000 \times 10^{11}$  molecules  
 C)  $4.000 \times 10^{10}$  molecules  
 D)  $3.000 \times 10^{10}$  molecules  
 E)  $7.000 \times 10^{10}$  molecules

9. Calculate the mass of carbon dioxide produced from  $1.0 \times 10^{-3}$  mol octane,  $\text{C}_8\text{H}_{18}$ , in the following reaction.



A)  $7.0 \times 10^{-3}$  g  
 B)  $1.000 \times 10^{-3}$  g  
 C)  $9.0 \times 10^{-3}$  g  
 D)  $1.200 \times 10^{-3}$  g  
 E)  $1.600 \times 10^{-3}$  g

10. Calculate the molecules of oxygen required to react with  $1.0 \times 10^{-3}$  mol sulfur in the following reaction.



A)  $1.000 \times 10^{10}$  molecules  $\text{O}_2$   
 B)  $1.000 \times 10^{11}$  molecules  $\text{O}_2$   
 C)  $4.000 \times 10^{10}$  molecules  $\text{O}_2$   
 D)  $3.000 \times 10^{10}$  molecules  $\text{O}_2$   
 E)  $7.000 \times 10^{10}$  molecules  $\text{O}_2$

# Multiple Choice Stoichiometry Test

**Katrin Zwirglmaier**

## Multiple Choice Stoichiometry Test:

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STEM based learning environments Highlighting a range of pivotal topics such as mobile games virtual labs and participatory simulations this publication is ideally designed for educators professionals academics and students seeking material on emerging educational technologies **The ETS Test Collection Catalog** Educational Testing Service. Test Collection, 1993

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