

Chemistry 2nd Semester Final Review KEY

This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

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Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

1. If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)

$$V_1 = 3.8 \text{ L}; P_1 = 2.71 \text{ atm}$$

$$V_2 = 1.47 \text{ L}; P_2 = ? \text{ atm}$$

Since all units are standard (L, atm), we do not need to convert the units before proceeding.

$$(P_1 V_1) / T_1 = (P_2 V_2) / T_2$$

T goes away (not mentioned in problem)

$$P_1 V_1 = P_2 V_2$$

Isolate for P2: $P_2 = (P_1 V_1) / V_2$

$$P_2 = (2.71 \text{ atm} \times 3.8 \text{ L}) / (1.47 \text{ L}) = 7.0 \text{ atm} \text{ (L cancelled and it's 2 sig figs)}$$

2. A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)

$$V_1 = 1.49 \text{ L}, T_1 = 34.75 \text{ degrees Celsius}$$

$$V_2 = ? \text{ L}, T_2 = 78.41 \text{ degrees Celsius}$$

We need to convert from Celsius to Kelvins before proceeding.

$$T_1 = 34.75 + 273 = 307.75 \text{ K}$$

$$T_2 = 78.41 + 273 = 351.41 \text{ K}$$

$$(P_1 V_1) / T_1 = (P_2 V_2) / T_2$$

Get rid of P

$$V_1 / T_1 = V_2 / T_2$$

Isolate V2: $V_2 = (V_1 / T_1) T_2$

$$V_2 = (1.49 \text{ L} / 307.75 \text{ K}) (351.41 \text{ K}) = 1.70 \text{ L}$$

3. What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?

The combined gas law won't work here. You only have one condition and nothing changes.

$PV = nRT$ will work. (n = moles and $R = 0.08026 \text{ L atm/K mol}$)

$P = 1.04 \text{ atm}$, $V = ? \text{ L}$, $n =$ OOPS we need to convert from grams to moles!, $T = 84.7 \text{ degrees Celsius} + 273 = 357.7 \text{ K}$

n : Hydrogen gas is diatomic!

$$8.47 \text{ g H}_2 \times (1 \text{ mol H}_2 / 2.02 \text{ g H}_2) = 4.193069307 \text{ mol H}_2$$

$$\text{Isolate V: } V = (nRT) / P = (4.193069307 \text{ mol H}_2 \times 0.08026 \text{ L atm/K mol} \times 357.7 \text{ K}) / 1.04 \text{ atm} = 118 \text{ L}$$

4. What volume is occupied by 56.75 g of oxygen gas at STP?

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L Towne

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Table of Contents 2nd Semester Final Review Guide Chemistry

1. Understanding the eBook 2nd Semester Final Review Guide Chemistry
 - The Rise of Digital Reading 2nd Semester Final Review Guide Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying 2nd Semester Final Review Guide Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2nd Semester Final Review Guide Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2nd Semester Final Review Guide Chemistry
 - Personalized Recommendations
 - 2nd Semester Final Review Guide Chemistry User Reviews and Ratings
 - 2nd Semester Final Review Guide Chemistry and Bestseller Lists

5. Accessing 2nd Semester Final Review Guide Chemistry Free and Paid eBooks
 - 2nd Semester Final Review Guide Chemistry Public Domain eBooks
 - 2nd Semester Final Review Guide Chemistry eBook Subscription Services
 - 2nd Semester Final Review Guide Chemistry Budget-Friendly Options
6. Navigating 2nd Semester Final Review Guide Chemistry eBook Formats
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 - 2nd Semester Final Review Guide Chemistry Compatibility with Devices
 - 2nd Semester Final Review Guide Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2nd Semester Final Review Guide Chemistry
 - Highlighting and Note-Taking 2nd Semester Final Review Guide Chemistry
 - Interactive Elements 2nd Semester Final Review Guide Chemistry
8. Staying Engaged with 2nd Semester Final Review Guide Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2nd Semester Final Review Guide Chemistry
9. Balancing eBooks and Physical Books 2nd Semester Final Review Guide Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2nd Semester Final Review Guide Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 2nd Semester Final Review Guide Chemistry
 - Setting Reading Goals 2nd Semester Final Review Guide Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2nd Semester Final Review Guide Chemistry
 - Fact-Checking eBook Content of 2nd Semester Final Review Guide Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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