

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentration from pH and hydroxide ion concentration from pOH.
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|-----------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Read the paragraph and use the information to determine if the sentence is true or false. Write true or false in the space provided. If the sentence is false, rewrite it so that it is a true statement.

1. When a substance has a pH of 1, it is an acid solution.
2. A substance with a pH of 10 is a base solution.
3. A substance with a pH of 7 is a neutral solution.
4. A substance with a pH of 10 is a base solution.
5. A substance with a pH of 10 is a base solution.
6. A substance with a pH of 10 is a base solution.
7. A substance with a pH of 10 is a base solution.
8. A substance with a pH of 10 is a base solution.
9. A substance with a pH of 10 is a base solution.
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18. A substance with a pH of 10 is a base solution.
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**Kevin T. Patton, Gary A.
Thibodeau, Andrew Hutton**



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