

- Structure, Hybridization, Resonance, Aromaticity
  - # of bonds- 1=sp (e.g. 3 bonds =  $sp^2$ )
  - Aromaticity=  $4n+2$  electrons
  - Strained ring causes greater heat of combustion
- Acids and Bases
  - Lewis acid= accepts electron pair
  - Higher percentage of s character means more acidic hydrogen
- Stereoisomerism
  - Constitutional Isomers
    - Isomers that have different atom connectivity
  - Stereoisomers
    - Same connectivity but different arrangement in space
    - Enantiomers
      - Stereoisomers that are nonsuperimposable mirror reflections
    - Diastereomers
      - Stereoisomers that are not mirror reflections of each other
  - Stereocenter=  $sp^3$  carbon bonded to four different substituents
  - Assigning R/S configuration
    - If lowest group is pointing away, determine normally
    - If lowest group is not pointing away, exchange with group pointing away and determine reverse
- Nucleophilic Substitutions and Eliminations
  - | Alkyl Halide | Strong Base | Weak Base                         | Poor Nucleophile                   |
|--------------|-------------|-----------------------------------|------------------------------------|
| 1            | $Sn2$       | $Sn2$                             | No rxn.                            |
| 2            | Mostly $E2$ | $Sn2$                             | $Sn1/E1$ in polar, protic solvents |
| 3            | $E2$        | $Sn1/E1$ in polar, protic solvent | $Sn1/E1$ in polar, protic solvents |
  - Relative strength of nucleophiles
    - Negatively charged nucleophiles react faster in substitution reactions than neutral nucleophiles
    - Nucleophilic strength is in the same order as base strength when the nucleophilic atoms are in the same period of the periodic table
    - Nucleophilic strength increases with increasing atom size when the nucleophilic atoms are in the same column of the periodic table
  - Leaving group ability
    - Increases with decreasing base strength
  - Base strength and structure

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**David Kirk**

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