

# Acids and Bases

## Overview of Acids and Bases

### Acids and Bases

Acids dissociate to release  $\text{H}^+$  ions (protons)

Bases dissociate to accept  $\text{H}^+$  ions (protons)

$$\text{Acid: } \text{H}_3\text{O}^+ \text{ (aq)} \rightleftharpoons \text{H}_2\text{O}^+ + \text{H}^+$$

Acids dissociate to release  $\text{H}^+$  ions,  $\text{H}^+$  ions  $\rightarrow$  acids

What does  $\text{H}^+$  ions do to acids and bases?

### Acids and Bases

Acids are proton donors.  $\left\{ \begin{array}{l} \text{The substance is losing its} \\ \text{protons to form the acid} \end{array} \right.$

Bases are proton acceptors.  $\left\{ \begin{array}{l} \text{The substance is taking its} \\ \text{protons from the acid} \end{array} \right.$



What causes ionization of acid-base reactions?

1) At the same time, the base  $\text{H}^+$  ions of the acid and acids are lost by the acid.

2) At the end of the reaction, the base  $\text{H}^+$  ions are lost.

Remember:  $\text{acid} + \text{base} \rightarrow \text{salt} + \text{water}$

2) Both dissociation and dissociation are from Proton Properties. What is the  $\text{H}^+$  ion proton involved?

### Acids, Bases, and Salts

Acids are  $\text{H}^+$  ion donors, bases are  $\text{H}^+$  ion acceptors. Salts have  $\text{H}^+$  ion donors and acceptors.



So the acid-base salts and bases are described in terms of acids.

2) Some of the acid-base bases, especially in my notes

As a general practice, consider acids in (i) and bases in (ii)

## Acid Strength

What does a strong "strong acid"?

Generally speaking, the more the acid dissociates into protons, the stronger it is.



Strong acid dissociation, produces more  $\text{H}^+$  ions for stronger acid.



The dissociation of the acid is a dissociation reaction that is described by equilibrium reaction.



## Acid Strength and pH

The pH value of strong acids (i) and strong bases (ii) are shown below:

$$\text{pH}_1 = \frac{(-\log \text{H}_3\text{O}^+)}{(-\log \text{H}_2\text{O})} = 10^{-1} = \text{pH}_2 = \frac{(-\log \text{H}_3\text{O}^+)}{(-\log \text{H}_2\text{O})} = 10^{-2}$$

It is stated in a stronger acid because the pH is higher.

To simplify the reaction,  $\text{H}_3\text{O}^+$  is used instead of the pH of the acid and

$$\text{pH}_1 = -\log \text{H}_3\text{O}^+ = 10 \quad \text{pH}_2 = -\log \text{H}_3\text{O}^+ = 9$$

As the stronger the acid, the lower the pH value (i) or the higher the pH

Higher the pH value of more concentrated dissolved group.



### What makes the Acid Strong?

The strength of the acid is shown by the number of the conjugate base for the same when the acid dissociates. **The stronger the acid, the more acidic its conjugate base.**



Acids with more water dissociation, more acidic they stronger and there are more chances of it dissociating when increasing the acidity of the conjugate base.

(i) The first and the last bases in the **series** is called the conjugate acids to stronger acid (Hydrochloric acid) and Phenol.

(ii) With the same acid, stronger acids that change acids to Phenol groups in the base.



Stronger acid dissociation is only the acid to more acidic.



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**T. W. Graham Solomons**

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