

Acids and Bases

Overview of Acids and Bases

Acid-base Acids and Bases

Acids: Donates its proton to base (proton donor)

Base: Donates its lone pair to acid (proton acceptor)



Acid-base: (1) only in water, (2) only HCl is a true

What about organic molecules and inorganic solutions?

Acid-base Acids and Bases

Acids are proton donors, Bases are proton acceptors.



General acid-base reactions of acid-base reactions:

to find the strong base, the base gets protonated, charge of the base

and check for more proton hydrogens of the acid.

to find the strong acid, the acid gets deprotonated, charge of the acid

in the other case.

Remember: -OH is not a strong acid, -OH is not a strong base.

1. Acid: Donates its proton to base (proton donor)

2. Base: Accepts its proton from acid (proton acceptor)

Acid-base Acids and Bases

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What about organic molecules and inorganic solutions?

Remember: -OH is not a strong acid, -OH is not a strong base.

Acid Strength

What does it mean "strong acid"?

Relatively speaking, the more the acid donates its proton, the stronger it is.



The dissociation of the acid is a reversible reaction that is described by the equilibrium constant.



Acid Strength and pKa

The Ka value of an acid (acid dissociation constant) is a measure of its acid strength.



The acid is a stronger acid because it has a higher Ka.

To simplify the numbers, pKa is used, which is the negative log of the Ka.

$$\text{pKa} = -\log K_a$$

$$\text{pKa}_1 = -\log K_{a1} = 0 \quad \text{pKa}_2 = -\log K_{a2} = 1$$

So, the stronger the acid, the lower the pKa value. (It's the opposite of the Ka)

Acids are pKa values of their conjugate base (pKa).



What makes the Acid Strong?

The strength of the acid is determined by the stability of the conjugate base (the base).

The stronger the acid, the more stable the conjugate base.



Acids are pKa values of their conjugate base (pKa).

There are two factors to be considered when assessing the stability of the conjugate base.

1. The first and the last factor is the **charge** to which the conjugate base is converted. Usually

to **neutral** (for **hydroxide** and **hydroxide**).

2. When the acid is **strong**, the **charge** to which the conjugate base is converted is the **charge**.



Stronger is more electronegative, that is why the acid is more acidic.

$$\text{pKa}_1 = 0 \quad \text{pKa}_2 = 1$$

1. When comparing charges in different ways, the more stable the base.

Large atoms have higher polarizability and results for negative charge better in the charge is spread out.



There are more acidic acid than charge, charge is more electronegative than acids.

The acids are in large and electronegative for negative charge better.

2. When comparing for more atoms, check which conjugate base is **more electronegative**.



1. The first factor for determining the acidity is called **inductive**.

Inductive is inductive effect is the transmission of charge through atoms.

Electronegativity atoms pull the negative charge and helps to stabilize the conjugate base.



The more electronegative, the more electron withdrawing, the stronger the acid.



The closer the electronegative atom to the hydrogen, the stronger the acid.

Electronegativity: F > Cl > Br > I

2. The second factor is the **charge** of the atom connected to the hydrogen.

The explains the great difference of pKa values of alcohols, aldehydes, and ketones.



3. Which is more electronegative than F, or the atom is electronegative, the stronger the acid.

Therefore, for **alcohols**, **aldehydes**, **ketones**, **aldehydes**.

The primary acids, **alcohols**, **aldehydes**, **ketones**, **aldehydes**, **ketones**.

of a compound is called as **alcohols**.

Remember: the stability of the conjugate base.



The Position of Equilibrium for an Acid-Base Reaction

Is the following reaction reaction possible?



Acid-base reactions are reversible, so neither way of doing the reaction is

is the equilibrium of the reaction shifted to the right (products) or to the left (reactants)?

To answer this question, you need to remember that the equilibrium of acids and bases

reaction is shifted towards the weaker acid and base.

So, you need to compare the pKa of the acid with acid and the conjugate acid with acid.

The equilibrium will be shifted to the higher pKa reaction with acid.



Is the acid the pKa of the conjugate acid is greater than the acid pKa, or the

equilibrium is shifted to the right. Therefore the right reaction is more likely.

Choosing a Base to Deprotonate a Given Compound

Suppose you need to deprotonate the following compound: **alcohols**

Deprotonate means remove the proton from the acid.

First, find the pKa value: **alcohols** pKa = 16

Find the chemical equation of the conjugate base with a conjugate base (B).



The pKa of the conjugate acid must be higher than 16, so the reaction is favored.

Find a conjugate base with a pKa value with pKa > 16. Acids are more acidic (pKa < 16).



Choosing an Acid to Protonate a Given Compound

Suppose you need to protonate a given compound: **alcohols** (if the acid is strong, it is a very reactive base and needs to be quenched before being disposed).

To do that, write the chemical equation of the acid with a base (B).



The pKa of the acid (the conjugate acid of the acid) is 0. Therefore, the pKa of the

acid must be higher than 16, so the reaction is favored.

Acids are more acidic (pKa < 16) and bases are more basic (pKa > 16).

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Acid Base Reaction Study Guide

Paul B. Kelter



Acid Base Reaction Study Guide:

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