

# Acids, Bases & pH Review - Honors Chemistry

Directions: For all multiple choice questions, highlight the correct answer(s). Answer short answer questions in the box provided.

## Part 1: Properties & Reactions of Acids & Bases

1. Complete the table below.

	ACIDS	BASES
Table		
pH range		
Conducting electricity		
Potassium		
Lithium		
Phenolphthalein		

2. An indicator:

- a. is the type of acid.
- b. is the type of base.
- c. has a neutral pH.
- d. changes color depending on the pH.

3. Lithium and phenolphthalein are indicators. What does this mean?

4. Acids and bases "dissociate" in water to form  $\text{H}^+$  (or  $\text{H}_3\text{O}^+$ ) and  $\text{OH}^-$  (or  $\text{H}_2\text{O}_2^-$ ). What does "dissociate" mean in this context?

5. Identify whether each substance is an acid or a base.

- a.  $\text{H}_2\text{O}^+$ :
- b.  $\text{HCl}$  ( $\text{HCl}_2$ ):
- c.  $\text{Al(OH}_3\text{)}_3$ :
- d.  $\text{NH}_3 + \text{H}_2\text{O} \rightarrow \text{NH}_4^+ + \text{OH}^-$ ;  $\text{NH}_3$  is weak.
- e.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{HSO}_4^-$ ;  $\text{H}_2\text{SO}_4$  is strong.
- f.  $\text{HNO}_3 \rightarrow \text{H}^+ + \text{NO}_3^-$ ;  $\text{HNO}_3$  is strong.

# Acid And Base Study Guide 2011

**S Marginson**

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## **Table of Contents Acid And Base Study Guide 2011**

1. Understanding the eBook Acid And Base Study Guide 2011
  - The Rise of Digital Reading Acid And Base Study Guide 2011
  - Advantages of eBooks Over Traditional Books
2. Identifying Acid And Base Study Guide 2011
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acid And Base Study Guide 2011
  - User-Friendly Interface
4. Exploring eBook Recommendations from Acid And Base Study Guide 2011
  - Personalized Recommendations
  - Acid And Base Study Guide 2011 User Reviews and Ratings
  - Acid And Base Study Guide 2011 and Bestseller Lists
5. Accessing Acid And Base Study Guide 2011 Free and Paid eBooks
  - Acid And Base Study Guide 2011 Public Domain eBooks
  - Acid And Base Study Guide 2011 eBook Subscription Services
  - Acid And Base Study Guide 2011 Budget-Friendly Options
6. Navigating Acid And Base Study Guide 2011 eBook Formats

- ePub, PDF, MOBI, and More
- Acid And Base Study Guide 2011 Compatibility with Devices
- Acid And Base Study Guide 2011 Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Acid And Base Study Guide 2011
- Highlighting and Note-Taking Acid And Base Study Guide 2011
- Interactive Elements Acid And Base Study Guide 2011

8. Staying Engaged with Acid And Base Study Guide 2011

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Acid And Base Study Guide 2011

9. Balancing eBooks and Physical Books Acid And Base Study Guide 2011

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Acid And Base Study Guide 2011

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Acid And Base Study Guide 2011

- Setting Reading Goals Acid And Base Study Guide 2011
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Acid And Base Study Guide 2011

- Fact-Checking eBook Content of Acid And Base Study Guide 2011
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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