

General Chemistry Reference Sheet

This reference sheet addresses some of the more peculiar pieces of information that need to be memorized in a general chemistry course. It also contains a simple set of essential formulas in chemistry with cautions, explanations, and general tips. This sheet is meant to be as concise as possible, and many informations in the textbook is left out in favor of cautions and tips. This sheet is, therefore, best used as a supplement to, not a replacement of, the textbook.

SI Fundamental Units

Mass	Kilogram (kg)
Length	Meter (m)
Time	Second (s)
Temperature	Kelvin (K)
Amount of substance	Mole (mol)
Electric current	Ampere (A)
Luminous intensity	Candela (cd)

Atomic Experiments and Models

J. J. Thomson	Discovered e^{-} ; Cathode ray
	Plum pudding model
R. A. Millikan	Measured charge of e^{-} ; Oil drop
H. Becquerel/M. Curie	Discovered radioactivity
E. Rutherford	Discovered α , β , and γ rays
	Discovered nucleus; Gold foil experiment
J. Chadwick	Discovered neutrons
N. Bohr	Bohr model (electron orbits)
Quantum mechanists	Quantum model

Polyatomic Ions

NH_4^{+}	ammonium	OH^{-}	hydroxide
CN^{-}	cyanide	$\text{C}_2\text{O}_4^{2-}$	oxalate
O_2^{2-}	peroxide	CNO^{-}	cyanate
HSO_3^{-}	hydrogen sul- fate	$\text{C}_2\text{H}_3\text{O}_2^{-}$	acetate
SCN^{-}	thiocyanate	NO_3^{-}	nitrate
SO_3^{2-}	sulfite	ClO_4^{-}	perchlorate
CO_3^{2-}	carbonate	ClO_3^{-}	chlorate
PO_4^{3-}	phosphate	ClO^{-}	hypochlorite
$\text{S}_2\text{O}_4^{2-}$	dithionite	HPO_4^{2-}	hydrogen phosphate
CrO_4^{2-}	chromate	H_2O^{+}	hydronium
$\text{Cr}_2\text{O}_7^{2-}$	dichromate	PO_3^{2-}	phosphite
MnO_4^{-}	permanganate	Hg_2^{2+}	mercury (I)
N_3^{-}	azide	C_4^{4-}	carbide
$\text{C}_4\text{H}_4\text{O}_6^{2-}$	tartarate	S_2^{2-}	disulfide
O_2^{-}	superoxide	AsO_3^{3-}	arsenite
PO_2^{3-}	hypophosphite	AsO_4^{3-}	arsenate
SiO_3^{2-}	silicate	$\text{P}_3\text{O}_9^{3-}$	pyrophosphate

Ionic Solubility Chart

Soluble	Exceptions
NO_3^{-}	None
$\text{CH}_3\text{COO}^{-}$	None
Cl^{-}	Ag^{+} , Hg_2^{2+} , Pb^{2+}
Br^{-}	Ag^{+} , Hg_2^{2+} , Pb^{2+}
I^{-}	Ag^{+} , Hg_2^{2+} , Pb^{2+}
SO_4^{2-}	Sr^{2+} , Ba^{2+} , Hg_2^{2+} , Pb^{2+}
Insoluble	Exceptions
S^{2-}	NH_4^{+} , alkali metal cations, Ca^{2+} , Sr^{2+} , Ba^{2+}
CO_3^{2-}	NH_4^{+} , alkali metal cations
PO_4^{3-}	NH_4^{+} , alkali metal cations
OH^{-}	NH_4^{+} , alkali metal cations, Ca^{2+} , Sr^{2+} , Ba^{2+}

Strong Acids and Bases

Strong acids and bases dissociate in water completely.

Strong Acids	Strong Bases	
HCl	HClO_4	Alkali metal hydroxides
HBr	HNO_3	Ca(OH)_2
HI	H_2SO_4	Sr(OH)_2
HClO_3	Ba(OH)_2	

Activity Series

Metals below H^{+} cannot react with acids to form H_2 . More active metals are better reducing agents.

From most active to least active:

Li^{+} , K^{+} , Ba^{2+} , Ca^{2+} , Na^{+} , Mg^{2+} , Al^{3+} , Mn^{2+} , Zn^{2+} , Cr^{3+} , Fe^{2+} , Co^{2+} , Ni^{2+} , Sn^{2+} , Pb^{2+} , H^{+} , Cu^{2+} , Ag^{+} , Hg_2^{2+} , Pt^{2+} , Au^{3+}

Flame Colors

Calcium	Brick red
Copper (I)	Blue
Copper (II)	Green or blue-green
Potassium	Lilac
Lithium	Dark red
Sodium	Bright yellow
Strontium	Red
Barium	Light green
Iron (III)	Gold
Cesium	Blue-Violet
Iodine	Blue
Lead	Blue
Rubidium	Red-Violet

Phase Changes

	From solid	From liquid	From gas
To solid	-	freezing	deposition
To liquid	melting	-	condensation
To gas	sublimation	evaporation	-

Solution Colors

Copper (II)	Blue
Nickel	Green
Permanganate	Purple
Chromate	Yellow
Dichromate	Orange
Iron (II)	Light blue
Iron (III)	Rusty yellow

Thermodynamic Laws

First Law: Energy cannot be created nor destroyed. It can only be transferred in the form of either heat or work.

Second Law: Any spontaneous reaction increases the entropy of the universe.

Third Law: An ideal solid crystal at 0 K has an entropy of 0.

Thermodynamic Formulas

Standard thermodynamic conditions	298 K; 1 atm; 1 M
Kinetic energy	$K = mv^2/2$
Electrostatic potential energy	$U_E = (k_e Q_1 Q_2)/d$
Internal energy	$\Delta E = q + w$
Enthalpy	$H = E + PV$
Specific heat	$s = q/(m \cdot \Delta T)$
Entropy in reversible reaction	$\Delta S_{\text{system}} = (\Delta H)/T$ $\Delta S_{\text{universe}} = -(\Delta H)/T$
Microstate-entropy relationship	$S = k \ln W$
Gibbs free energy	$G = H - TS$
Gibbs free energy change	$\Delta G = \Delta H - T \Delta S$ $\Delta G = \Delta G^{\circ} + RT \ln Q$ $\Delta H_{\text{total}} = \sum \Delta H_i$
Hess's Law	

Constants

Boltzmann's constant	$k_B = 1.381 \times 10^{-23} \text{ m}^2\text{kg} \cdot \text{s}^{-2}\text{K}^{-1}$
Coulomb's constant $k_C = 1/(4\pi\epsilon_0)$	$= 8.988 \times 10^9 \text{ J} \cdot \text{m}^2/\text{C}^2$
Avogadro's number	$N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$
Faraday's constant	$F = 9.648 \times 10^4 \text{ C/mol}$
Planck's constant	$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$
Ideal gas constants	$R = 0.0821 \text{ (L} \cdot \text{atm)/(mol} \cdot \text{K)}$ $R = 8.314 \text{ J/(mol} \cdot \text{K)}$
Vacuum permittivity $\epsilon_0 = 1/(\mu_0 c^2)$	$= 8.854 \times 10^{-12} \text{ F/m}$
Vacuum permeability	$\mu_0 = 1.257 \times 10^{-6} \text{ N}\cdot\text{A}^{-2}$
Atomic mass	$1 \text{ amu} = 1.661 \times 10^{-24} \text{ g}$
Electron charge	$e = 1.602 \times 10^{-19} \text{ C}$
Electronvolt	$1 \text{ eV} = 1.602 \times 10^{-19} \text{ J}$
Atmospheric pressure	$1 \text{ atm} = 1.013 \times 10^5 \text{ Pa}$
Absolute zero	$0 \text{ K} = -273.15 \text{ }^{\circ}\text{C}$
Speed of light in vacuum	$c = 2.998 \times 10^8 \text{ m/s}$

Quantum Mechanical Formulas

Energy of a quantum	$E = hf$
Wavelength-frequency relationship	$c = \nu \cdot \lambda$
Probability distribution	$P_V = \iint \psi(x,y,z) ^2 dxdydz$

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Gershon J. Shugar, Jack T. Ballinger



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Table of Contents Acs Study Guide General Chemistry Formula Sheet

1. Understanding the eBook Acs Study Guide General Chemistry Formula Sheet
 - The Rise of Digital Reading Acs Study Guide General Chemistry Formula Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Acs Study Guide General Chemistry Formula Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acs Study Guide General Chemistry Formula Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acs Study Guide General Chemistry Formula Sheet
 - Personalized Recommendations
 - Acs Study Guide General Chemistry Formula Sheet User Reviews and Ratings
 - Acs Study Guide General Chemistry Formula Sheet and Bestseller Lists

5. Accessing Acs Study Guide General Chemistry Formula Sheet Free and Paid eBooks
 - Acs Study Guide General Chemistry Formula Sheet Public Domain eBooks
 - Acs Study Guide General Chemistry Formula Sheet eBook Subscription Services
 - Acs Study Guide General Chemistry Formula Sheet Budget-Friendly Options
6. Navigating Acs Study Guide General Chemistry Formula Sheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Acs Study Guide General Chemistry Formula Sheet Compatibility with Devices
 - Acs Study Guide General Chemistry Formula Sheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acs Study Guide General Chemistry Formula Sheet
 - Highlighting and Note-Taking Acs Study Guide General Chemistry Formula Sheet
 - Interactive Elements Acs Study Guide General Chemistry Formula Sheet
8. Staying Engaged with Acs Study Guide General Chemistry Formula Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acs Study Guide General Chemistry Formula Sheet
9. Balancing eBooks and Physical Books Acs Study Guide General Chemistry Formula Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acs Study Guide General Chemistry Formula Sheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Acs Study Guide General Chemistry Formula Sheet
 - Setting Reading Goals Acs Study Guide General Chemistry Formula Sheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acs Study Guide General Chemistry Formula Sheet
 - Fact-Checking eBook Content of Acs Study Guide General Chemistry Formula Sheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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