
ACS Biochemistry Practice Final, FGCU, Beharry, Coticone With Complete Solution 2024

1. A weak acid, HA, has a pKa of 4.11.

- a) At pH 7.11, what will the ratio of conjugate base to acid be? (Henderson-Hasselbach equation: $\text{pH} = \text{pKa} + \log \frac{[\text{A}^-]}{[\text{HA}]}$)
 - b) Based on answer to "a," is there more HA or A- at that pH?
 - c) At pH 2.11, what will the ratio of conjugate base to acid be?
 - d) Based on answer to "c," is there more HA or A- at that pH?
 - e) At pH 4.11, what will the ratio of conjugate base to acid be?
 - f) Fill in the blank: if pH is greater than the pKa of a weak acid, there will be ____ A- relative to HA. If the pH is less than the pKa of a weak acid, there will be ____ A- relative to HA. If the pH=pKa, there will be ____ of A- and HA. - a) $7.11 = 4.11 + \log(x)$
 $3 = \log(x)$
 $1000 = x$
 - b) A-
 - c) $2.11 = 4.11 + \log(x)$
 $-2 = \log(x)$
 $0.01 = x$
 - d) HA
 - e) 1
 - f) more; less; the same amount
- 2.
- a) Which amino acids would be charged at pH 7?
 - b) which amino acids from "a" would be negatively charged?
 - c) which amino acids from "a" would be positively charged? - a) Asp, Glu, Arg, His, Lys
 - b) Asp, Glu
 - c) Arg, His, Lys
3. Which of the following sequences is most likely to be in the interior core of a globular protein? Why?

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