
ACS Biochemistry Practice Final, FGCU, Beharry, Coticone With Complete Solution 2024

1. A weak acid, HA, has a pKa of 4.11.

a) At pH 7.11, what will the ratio of conjugate base to acid be? (Henderson-Hasselbach equation: $\text{pH} = \text{pK}_a + \log \frac{[\text{A}^-]}{[\text{HA}]}$)

b) Based on answer to "a," is there more HA or A⁻ at that pH?

c) At pH 2.11, what will the ratio of conjugate base to acid be?

d) Based on answer to "c," is there more HA or A⁻ at that pH?

e) At pH 4.11, what will the ratio of conjugate base to acid be?

f) Fill in the blank: if pH is greater than the pKa of a weak acid, there will be ____ A⁻ relative to HA. If the pH is less than the pKa of a weak acid, there will be ____ A⁻ relative to HA. If the pH=pKa, there will be ____ of A⁻ and HA. - a) $7.11 = 4.11 + \log[x]$

$$3 = \log[x]$$

$$1000 = x$$

b) A⁻

$$c) 2.11 = 4.11 + \log[x]$$

$$-2 = \log[x]$$

$$0.01 = x$$

d) HA

e) 1

f) more; less; the same amount

2.

a) Which amino acids would be charged at pH 7?

b) which amino acids from "a" would be negatively charged?

c) which amino acids from "a" would be positively charged? - a) Asp, Glu, Arg, His,

Lys

b) Asp, Glu

c) Arg, His, Lys

3. Which of the following sequences is most likely to be in the interior core of a globular protein? Why?

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