

ACS Practice Test 1

Acids & Bases:

1. All are examples of Lewis acid–base reactions *except*

- (A) $\text{Cu}^{2+}(\text{aq}) + 4\text{NH}_3(\text{aq}) \rightleftharpoons [\text{Cu}(\text{NH}_3)_4]^{2+}(\text{aq})$
 (B) $\text{HCl}(\text{g}) + \text{NH}_3(\text{g}) \rightarrow \text{NH}_4\text{Cl}(\text{s})$
 (C) $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightleftharpoons \text{H}_2\text{O}(\text{l})$
 (D) $2\text{Na}(\text{s}) + \text{Cl}_2(\text{g}) \rightarrow 2\text{NaCl}(\text{s})$

2. According to the Lewis definition, an acid is a species

- (A) having a hydrogen ion. (B) donating a pair of electrons.
 (C) accepting a pair of electrons. (D) accepting a hydrogen ion.

3. All are potential Lewis bases *except*

- (A) NH_3 (B) H_2O (C) CH_4 (D) CN^-

4. Which hydrated metal ion is most acidic under conditions of equal molar concentration in water?

- (A) Al^{3+} (B) Ba^{2+} (C) K^+ (D) Zn^{2+} (E) Ag^+

5. Which element in Group 5A forms the most basic oxide?

- (A) N (B) P (C) As (D) Sb (E) Bi

Atomic Structure:

6. In what respect does an atom of magnesium differ from a magnesium ion (Mg^{2+})?

- (A) The ion has a more stable electronic arrangement than the atom.
 (B) The positive charge on the nucleus of the ion is two units greater than the nuclear charge on the atom.
 (C) The ion has two more protons than the atom.
 (D) The ion will react more readily with free Cl_2 .
 (E) The ion has two more planetary electrons than the atom.

7. A certain atom has the symbol $^{148}_{75}\text{Px}$. What does this symbol tell about an atom of the element?

- (A) It has 148 electrons. (B) It has a valence of 4.
 (C) It has 221 protons in the nucleus. (D) It has 75 neutrons in the nucleus.
 (E) Its atomic mass is 73 u.

8. The atomic mass of an element is 32.07 u and its atomic number is 16. The number of protons contained in the nucleus of an atom of this element is

- (A) 16 (B) 31 (C) 32 (D) 48

9. The atomic mass of an element is 32.07 u and its atomic number is 16. The element forms a simple ion. The ion will most likely have a charge of

- (A) 1– (B) 2– (C) 3+ (D) 1+

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1. Understanding the eBook Acs Practice Test 3 Answers
 - The Rise of Digital Reading Acs Practice Test 3 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Acs Practice Test 3 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acs Practice Test 3 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acs Practice Test 3 Answers
 - Personalized Recommendations

- Acs Practice Test 3 Answers User Reviews and Ratings
- Acs Practice Test 3 Answers and Bestseller Lists
- 5. Accessing Acs Practice Test 3 Answers Free and Paid eBooks
 - Acs Practice Test 3 Answers Public Domain eBooks
 - Acs Practice Test 3 Answers eBook Subscription Services
 - Acs Practice Test 3 Answers Budget-Friendly Options
- 6. Navigating Acs Practice Test 3 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Acs Practice Test 3 Answers Compatibility with Devices
 - Acs Practice Test 3 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acs Practice Test 3 Answers
 - Highlighting and Note-Taking Acs Practice Test 3 Answers
 - Interactive Elements Acs Practice Test 3 Answers
- 8. Staying Engaged with Acs Practice Test 3 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acs Practice Test 3 Answers
- 9. Balancing eBooks and Physical Books Acs Practice Test 3 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acs Practice Test 3 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acs Practice Test 3 Answers
 - Setting Reading Goals Acs Practice Test 3 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acs Practice Test 3 Answers
 - Fact-Checking eBook Content of Acs Practice Test 3 Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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