



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

Key Vocabulary

| | | |
|---|---|---|
| <ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base | <ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base | <ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base |
|---|---|---|

Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

$$\text{pH} = -\log_{10}[\text{H}^+]$$

Practice Calculations

Use the following equations to calculate the pH of a solution with a hydrogen ion concentration of $1.0 \times 10^{-5} \text{ M}$.

$\text{pH} = -\log_{10}[\text{H}^+]$ $\text{pH} = -\log_{10}[\text{H}^+]$
concentration concentration

| Concentration | pH |
|---------------------------------|-------|
| $1.0 \times 10^{-5} \text{ M}$ | 5.00 |
| $1.0 \times 10^{-6} \text{ M}$ | 6.00 |
| $1.0 \times 10^{-7} \text{ M}$ | 7.00 |
| $1.0 \times 10^{-8} \text{ M}$ | 8.00 |
| $1.0 \times 10^{-9} \text{ M}$ | 9.00 |
| $1.0 \times 10^{-10} \text{ M}$ | 10.00 |
| $1.0 \times 10^{-11} \text{ M}$ | 11.00 |
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