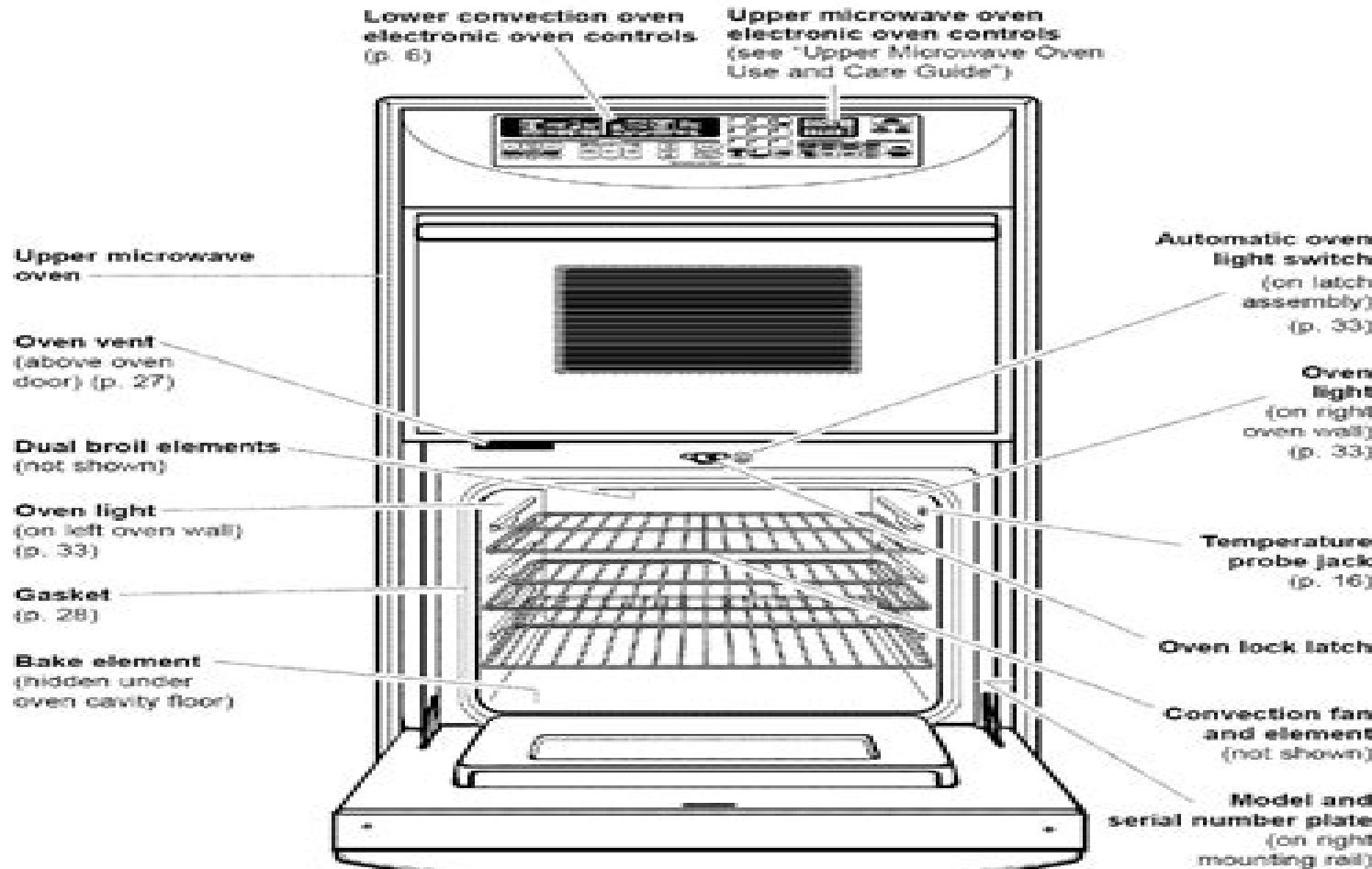


Getting to Know Your Electric Convection Oven

Oven parts and features



Operating Instructions Microwave Convection Oven

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Sebastian Brünink

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2 $\text{CO}_2 \text{ g} \rightarrow \text{H}_2 \text{O g}$ $\Delta h = 5 \text{ kJ}$
 $\text{C}_2\text{H}_2 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g}$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2 \text{ g} \rightarrow \frac{1}{2}\text{O}_2 \text{ g}$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} \rightarrow \text{C}_2\text{H}_2 \text{ g} + \text{H}_2 \text{ g}$ $\Delta h = 283 \text{ kJ}$
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web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kj in your final answer multiply the mass by the change in temperature and the c given trial 1 use hess s law and the following equations and Δh values to determine the heat of reaction for the reaction $c 2 h 4$

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web thus for reaction 9 2 in which 1 mole of $h3o$ and 1 mole of oh combine to form 2 moles of $h2o$ h 55 8 kj as shown note that the enthalpy change h hproducts hreactants is positive if heat is absorbed that is if hproducts

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web hess s law enthalpy of reaction and hess s law abstract the goal of this lab is to perform experiments and practice hess s law by determining enthalpy change in chemical reactions in this lab three experiments were conducted in order to test hess s law the experiments required observing and collecting data on three different reactions

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic

reactions 1 naoh s na aq oh aq h 1 2 naoh s h aq cl aq h 2 o na aq cl aq h 2 3