

Unit 8: Practice

1. The balanced equation $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$ tells us that 4.0 mol H_2

- A) reacts with 2.0 mol P_4 .
- B) produces 8.0 mol PH_3 .
- C) cannot react with phosphorus.
- D) produces 2.7 mol PH_3 .
- E) reacts with 4.0 mol P_4 .

2. A 1.0-mol sample of PCl_5 was decomposed according to the equation



How many moles of Cl_2 are formed assuming 100% yield?

- A) 0.5 mol
- B) 1.0 mol
- C) 1.2 mol
- D) 0.7 mol
- E) 2.0 mol

3. The rusting of iron is represented by the equation $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$. If you have a 1.04-mol sample of iron, how many moles of Fe_2O_3 will there be when the iron has rusted completely?

- A) 0.527 mol
- B) 0.790 mol
- C) 1.05 mol
- D) 1.58 mol
- E) 2.07 mol

4. Consider the following reaction, where X represents an unknown element.

$$5\text{X}_2(\text{s}) + 2\text{H}_2\text{O}_2(\text{l}) \rightarrow \text{H}_2\text{X}_2(\text{l}) + 2\text{XO}_2(\text{g})$$

If 175 g of X react completely with hydrogen peroxide to produce 1.43 mol of H_2X_2 , what is the identity of X?

- A) Ge
- B) Mg
- C) Se
- D) Sn
- E) C

5. Refer to the following equation:



How many molecules of NO_2 are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08
- B) 3.64×10^{23}
- C) 1.82×10^{23}
- D) 128
- E) none of these

6. In the reaction $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$, how many moles of ammonia would be produced from 1.08 mol of hydrogen and excess nitrogen?

- A) 1.42 mol
- B) 2.27 mol
- C) 0.727 mol
- D) 2.18 mol
- E) 0.543 mol

7. Refer to the following unbalanced equation:

$$\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

What mass of oxygen (O_2) is required to react completely with 13.1 g of C_4H_{10} ?

- A) 2.98×10^2 g
- B) 23.1 g
- C) 4.86 g
- D) 49.2 g
- E) 0.152 g

8. How many molecules of carbon dioxide will be formed if 7.07 g of propane is burned in the following reaction?



- A) 1.37×10^{23} molecules
- B) 1.03×10^{23} molecules
- C) 4.14×10^{23} molecules
- D) 3.10×10^{23} molecules
- E) 5.17×10^{23} molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane, C_8H_{18} , in the following reaction:



- A) 79.2 g CO_2
- B) 158 g CO_2
- C) 99.1 g CO_2
- D) 1.238 g CO_2
- E) 1.400 g CO_2

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction:



- A) 1.03×10^{23} molecules O_2
- B) 2.28×10^{23} molecules O_2
- C) 1.71×10^{23} molecules O_2
- D) 3.42×10^{23} molecules O_2
- E) 5.13×10^{23} molecules O_2

Multiple Choice Stoichiometry Test

Roger Slade, Maureen Slade



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