

Chemical Reactions:

1. The following reaction $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ is given. What is the ΔH°_f for the reaction?

A) greater than 2.00 kcal/mol
 B) greater than 1.00 kcal/mol
 C) less than 1.00 kcal/mol
 D) less than 2.00 kcal/mol
 E) greater than 4.00 kcal/mol

2. A 1.0-mol sample of PCl_3 reacts with excess oxygen according to the equation: $2\text{PCl}_3(\text{g}) \rightarrow 2\text{PCl}_5(\text{g}) + \text{O}_2(\text{g})$. Below, which reaction will yield the greatest amount of O_2 after PCl_3 is removed?

A) 0.0 mol
 B) 1.0 mol
 C) 1.2 mol
 D) 1.7 mol
 E) 2.0 mol

3. The reaction of NO_2 is step characterized by the equation: $4\text{NO}_2 \rightarrow 3\text{O}_2 + 2\text{N}_2\text{O}_5$. If 1.0 mol of NO_2 is added to 1.0 mol of N_2O_5 and the reaction is allowed to proceed, how will the reaction rate change?

A) 1.000 mol/d
 B) 0.750 mol/d
 C) 1.000 mol/d
 D) 1.50 mol/d
 E) 2.00 mol/d

4. Calculated are the following reactions, or from the experiments, are unknowns, calculated.

$\text{H}_2\text{O}_2(\text{l}) + \text{H}_2\text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O}_2(\text{l}) = 300\text{J/g}$

$\text{H}_2 + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$ (enthalpy of difference, ΔH°_f for production of 1.00 mol of H_2O_2 , which is the difference of 200)

A) O_2
 B) H_2
 C) H_2
 D) O_2
 E) C

5. Refer to the following reaction:

$\text{NO}_2(\text{g}) + \text{NO}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + \text{O}_2(\text{g})$

Major steady-state levels of NO_2 are given below. In which behavior is each?

A) 1.2×10^{-4}
 B) 1.25×10^{-4}
 C) 1.30×10^{-4}
 D) 1.35×10^{-4}
 E) Insufficient information

6. The following reaction $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ is given. Below, which reaction will yield the greatest amount of O_2 after 1.00 mol of hydrogen and oxygen react?

A) 1.00 mol
 B) 1.20 mol
 C) 1.25 mol
 D) 1.30 mol
 E) 1.35 mol

7. Refer to the following reaction:

$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CH}_3\text{OH} + \text{H}_2\text{O}$

What mass of oxygen, O_2 , is required to react completely with 1.00 g of C_2H_6 ?

A) 1.00×10^{-2} g
 B) 1.00×10^{-1} g
 C) 1.00×10^{-3} g
 D) 1.00×10^{-4} g
 E) 1.00×10^{-5} g

8. Refer reaction neutralization and synthesis of ammonia. In the following reaction, which is the limiting reactant?



A) 1.00 $\times 10^{-2}$ mol/deciliter
 B) 1.00 $\times 10^{-3}$ mol/deciliter
 C) 1.00×10^{-4} mol/deciliter
 D) 1.00×10^{-5} mol/deciliter
 E) 1.00×10^{-6} mol/deciliter

9. Calculate the mass of carbon dioxide produced from 1.00 g of octane, C_8H_{18} , in the following reaction:



A) 1.75×10^{-2} g
 B) 1.75×10^{-1} g
 C) 1.75×10^{-3} g
 D) 1.75×10^{-4} g
 E) 1.75×10^{-5} g

10. Calculate the mass of oxygen required to react with 1.00 g of sulfur in the following reaction:



A) 1.00×10^{-2} mol/deciliter O_2
 B) 1.00×10^{-3} mol/deciliter O_2
 C) 1.00×10^{-4} mol/deciliter O_2
 D) 1.00×10^{-5} mol/deciliter O_2
 E) 1.00×10^{-6} mol/deciliter O_2

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Journal of Chemical Education, 1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers *Tests* Richard C. Sweetland, Daniel J. Keyser, 1983 Descriptions of over 3000 tests in English intended as a guide for psychologists educators and other personnel who need test information to meet their assessment needs Topical arrangement under 3 main sections of psychology education and business Each entry gives test name associated personal names intended age group purpose description time range scoring cost and publisher Indexes by test titles authors publishers visually impaired tests and scoring services

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