



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base in terms of the Arrhenius theory.
- Define strong and weak acids.

Definitions

- Hydrogen ion: H^+
- Acid: a substance that dissociates in water to produce hydrogen ions.
- Base: a substance that dissociates in water to produce hydroxide ions.

Key Equations

- $\text{pH} = -\log[\text{H}^+]$
- $\text{pH} = \log[\text{OH}^-]$

$$\text{pH} = -\log[\text{H}^+]$$

Practice Calculations

Given the following values of hydrogen ion concentration, calculate the pH of each solution.

Hydrogen ion concentration: $1.0 \times 10^{-3} \text{ M}$ Hydrogen ion concentration: $1.0 \times 10^{-10} \text{ M}$

Hydrogen ion concentration: $1.0 \times 10^{-2} \text{ M}$ Hydrogen ion concentration: $1.0 \times 10^{-11} \text{ M}$

Hydrogen ion concentration: $1.0 \times 10^{-4} \text{ M}$ Hydrogen ion concentration: $1.0 \times 10^{-12} \text{ M}$

Answers

- 1.0
- 9.0
- 2.0
- 10.0
- 4.0
- 11.0
- 6.0
- 7.0
- 8.0
- 9.0
- 10.0
- 11.0
- 12.0

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Barbara Janson Cohen, Kerry L. Hull

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