

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. pH | |

Key Equation

- | | |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constants indicated that are associated in pH-related calculations with hydrogen ions and hydroxide ions.

1. What substance(s) is/are _____ in water? Hydrogen ion(s) and hydroxide ion(s).
2. The concentration of these species present in water is equal to _____ mol/L.
3. The product of their concentrations is _____ at 25°C.
4. The concentration of these species is _____ mol/L.
5. The _____ of these species is _____ mol/L.
6. The _____ of these species is _____ mol/L.
7. The _____ of these species is _____ mol/L.
8. The _____ of these species is _____ mol/L.
9. The _____ of these species is _____ mol/L.
10. The _____ of these species is _____ mol/L.
11. The _____ of these species is _____ mol/L.
12. The _____ of these species is _____ mol/L.
13. The _____ of these species is _____ mol/L.
14. The _____ of these species is _____ mol/L.
15. The _____ of these species is _____ mol/L.
16. The _____ of these species is _____ mol/L.
17. The _____ of these species is _____ mol/L.
18. The _____ of these species is _____ mol/L.
19. The _____ of these species is _____ mol/L.
20. The _____ of these species is _____ mol/L.

1. K_w
2. 1.0×10^{-14}
3. 1.0×10^{-7}
4. 1.0×10^{-7}
5. 1.0×10^{-7}
6. 1.0×10^{-7}
7. 1.0×10^{-7}
8. 1.0×10^{-7}
9. 1.0×10^{-7}
10. 1.0×10^{-7}
11. 1.0×10^{-7}
12. 1.0×10^{-7}
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20. 1.0×10^{-7}

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