

# American Chemical Society (A.C.S.) General Chemistry Exam

75 multiple choice questions, 110 minutes

Bring a scantron, calculator and pencil(s)

- I. **States of Matter**
  - Ideal Gas Law:  $PV = nRT$ ,  $PM = dRT$ , deviations from ideal (high P, low T)
  - Ionic Solids: sc, fcc, bcc
  - Phase Diagrams: Pressure versus Temperature
  - Intermolecular Forces: ion-dipole, hydrogen bonding, etc.
- II. **Stoichiometry and Thermochemistry**
  - Hess' Law & Enthalpy: "standard states" have  $\Delta H = 0$
  - Percent Composition: %C, %H, etc.
  - Limiting Reactant(s)
  - Stoichiometry
- III. **Atomic Structure and Periodicity**
  - Oxidation States of "Main group" elements (Li, Be, B, etc.)
  - Oxidation Numbers
  - Atomic orbital configurations & relation to groups on periodic table ( $ns^2np^5$  = halogens)
  - Periodic Trends: Electronegativity, atomic radii
  - Nuclear chemistry: mass number, atomic number
- IV. **Molecular Structure**
  - Modes of bonding: ionic, covalent (both polar and nonpolar)
  - Lewis Dot, VSEPR, polarity, bond angles, resonance structures
- V. **Solutions**
  - Molarity: mol solute / L solution; also calculating molarity,  $C_1V_1 = C_2V_2$
  - Net ionic equations
  - Percent dissociation in water (electrolyte strength)
  - Molality: mol solute / kg solvent
  - Freezing Point Depression  $\Delta T = k_m$
- VI. **Acids, Bases and Ionic Equilibria**
  - Know the "inert" acid/base ions:  $Cl^-$ ,  $NO_3^-$ ,  $Na^+$ , etc.
  - Know general reactions for W/A + S/B, W/B + S/A, S/B + S/A
  - Equivalence point, finding unknown concentrations
  - Buffer: what makes a buffer, what a buffer is, etc.
  - pH: calculate pH of W/B + S/A, S/B + W/A, pure W/B, pure W/A
- VII. **Chemical Equilibrium - Molecular**
  - General equilibrium expressions: know how to find K
  - Precipitation reactions
  - Effect of heat upon endothermic / exothermic reactions
- VIII. **Kinetics**
  - Rate of disappearance, rate of formation; Catalysis in mechanisms
  - Rate law and half-life from experimental data; Energy of activation
- IX. **Thermodynamics**
  - Enthalpy, entropy, free energy
  - Relation of free energy to equilibrium constants
- X. **Electrochemistry and Redox**
  - E values for electrochemical cells; E values for electrolysis cells
  - Balancing acidic redox reactions
  - Best reducing agents, oxidizing agents; definitions of oxidized and reduced
- XI. **Descriptive Chemistry**
  - Nuclear chemistry: particle(s) emitted which result in increase of atomic number
  - Elemental reactivity: with water, oxygen, bases, acid, etc.
  - Organic chemistry: know alkenes, alkanes, etc.; also isomers
  - Coordination chemistry: coordination numbers around central metal
  - Ionization energy periodic trends
  - General nomenclature
- XII. **Laboratory**
  - Effect of improper drying on mass measurements
  - Methods for finding  $[Cl^-]$

# American Chemical Society Study Guide 201

**Albert A Gayle**



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*Readers' Guide to Periodical Literature* Anna Lorraine Guthrie, Bertha Tannehill, Neltje Marie Tannehill Shimer, 1919  
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