

American Chemical Society (A.C.S.) General Chemistry Exam

73 multiple choice questions, 110 minutes
Bring a scanner, calculator and pencils!

I. States of Matter

Ideal Gas Law: $PV = nRT$, $PV = \rho RT$, deviations from ideal (high P, low T)
Ionic Solids: sc, fcc, bcc

Phase Diagrams: Pressure versus Temperature
Intermolecular Forces: ion-dipole, hydrogen bonding, etc.

II. Stoichiometry and Thermochemistry

Hess' Law & Enthalpy: "standard states" have $\Delta H = 0$
Percent Composition: %C, %H, etc.
Limiting Reactant(s)

III. Atomic Structure and Periodicity

Oxidation States of "Main group" elements (Li, Be, B, etc.)
Oxidation Numbers
Atomic orbital configurations & relation to groups on periodic table (ns²np⁶ = halogens)
Periodic Trends: Electronegativity, atomic radius
Nuclear chemistry: mass number, atomic number

IV. Molecular Structure

Modes of bonding: ionic, covalent (both polar and nonpolar)
Lewis Dot, VSEPR, polarity, bond angles, resonance structures

V. Solutions

Molarity: mol solute / L solution; also calculating molarity, $C_1V_1 = C_2V_2$
Net ionic equations
Percent dissociation in water (electrolyte strength)
Molality: mol solute / kg solvent
Freezing Point Depression $\Delta T = k_m$

VI. Acids, Bases and Ionic Equilibria

Know the "inert" acid/base ions: Cl^- , NO_3^- , Na^+ , etc.
Know general reactions for $WA + SB$, $WB + SA$, $SB + SA$
Equivalence point, finding unknown concentrations
Buffer: what makes a buffer, what a buffer is, etc.

pH: calculate pH of $WB + SA$, $SB + WA$, pure WB , pure WA

VII. Chemical Equilibrium - Molecular

General equilibrium expressions: know how to find K
Precipitation reactions
Effect of heat upon endothermic / exothermic reactions

VIII. Kinetics

Rate of disappearance, rate of formation; Catalysis in mechanisms
Rate law and half-life from experimental data; Energy of activation

IX. Thermodynamics

Enthalpy, entropy, free energy
Relation of free energy to equilibrium constants

X. Electrochemistry and Redox

E values for electrochemical cells; E values for electrolysis cells
Balancing acidic redox reactions
Best reducing agents, oxidizing agents; definitions of oxidized and reduced

XI. Descriptive Chemistry

Nuclear chemistry: particle(s) emitted which result in increase of atomic number
Elemental reactivity: with water, oxygen, bases, acid, etc.
Organic chemistry: know alkenes, alkanes, etc.; also isomers
Coordination chemistry: coordination numbers around central metal
Ionization energy periodic trends
General nomenclature

XII. Laboratory

Effect of improper drying on mass measurements
Methods for finding $[Cl^-]$

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Albert A Gayle

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Journal of the American Chemical Society American Chemical Society,1925 Proceedings of the Society are included in v 1 59 1879 1937 **ACS Monograph** American Chemical Society,1929 **Acid Precipitation** ,1992 **American**

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