

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentration from pH and hydroxide ion concentration from pOH.
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. pOH | 6. acid solution |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the hydrogen ion concentration or autoionization constant, if the missing quantity that is indicated in the partial equation below. Complete each equation as needed.

1. What substance has a pH of 10.0? Hydrogen ion concentration is 10^{-10} M. The autoionization of water which produces an OH^- ion and a H^+ ion is 10^{-14} M.
2. Hydrochloric acid has a pH of 1.0. The H^+ ion concentration is 10^{-1} M. The autoionization of water which produces an OH^- ion and a H^+ ion is 10^{-14} M.
3. The H^+ ion concentration of a solution is 10^{-12} M. The OH^- ion concentration is 10^{-2} M. The autoionization of water which produces an OH^- ion and a H^+ ion is 10^{-14} M.
4. The H^+ ion concentration of a solution is 10^{-12} M. The OH^- ion concentration is 10^{-2} M. The autoionization of water which produces an OH^- ion and a H^+ ion is 10^{-14} M.
5. The H^+ ion concentration of a solution is 10^{-12} M. The OH^- ion concentration is 10^{-2} M. The autoionization of water which produces an OH^- ion and a H^+ ion is 10^{-14} M.

6. $\text{pH} = 10.0$
7. $\text{pH} = 10.0$
8. $\text{pH} = 10.0$
9. $\text{pH} = 10.0$
10. $\text{pH} = 10.0$
11. $\text{pH} = 10.0$
12. $\text{pH} = 10.0$
13. $\text{pH} = 10.0$
14. $\text{pH} = 10.0$
15. $\text{pH} = 10.0$
16. $\text{pH} = 10.0$
17. $\text{pH} = 10.0$
18. $\text{pH} = 10.0$

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Arturo Cuomo



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Table of Contents 19 2 Section Review Hydrogen Ions And Acidity

1. Understanding the eBook 19 2 Section Review Hydrogen Ions And Acidity
 - The Rise of Digital Reading 19 2 Section Review Hydrogen Ions And Acidity
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Section Review Hydrogen Ions And Acidity
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Section Review Hydrogen Ions And Acidity
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Section Review Hydrogen Ions And Acidity

- Personalized Recommendations
- 19 2 Section Review Hydrogen Ions And Acidity User Reviews and Ratings
- 19 2 Section Review Hydrogen Ions And Acidity and Bestseller Lists
- 5. Accessing 19 2 Section Review Hydrogen Ions And Acidity Free and Paid eBooks
 - 19 2 Section Review Hydrogen Ions And Acidity Public Domain eBooks
 - 19 2 Section Review Hydrogen Ions And Acidity eBook Subscription Services
 - 19 2 Section Review Hydrogen Ions And Acidity Budget-Friendly Options
- 6. Navigating 19 2 Section Review Hydrogen Ions And Acidity eBook Formats
 - ePub, PDF, MOBI, and More
 - 19 2 Section Review Hydrogen Ions And Acidity Compatibility with Devices
 - 19 2 Section Review Hydrogen Ions And Acidity Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 2 Section Review Hydrogen Ions And Acidity
 - Highlighting and Note-Taking 19 2 Section Review Hydrogen Ions And Acidity
 - Interactive Elements 19 2 Section Review Hydrogen Ions And Acidity
- 8. Staying Engaged with 19 2 Section Review Hydrogen Ions And Acidity
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Section Review Hydrogen Ions And Acidity
- 9. Balancing eBooks and Physical Books 19 2 Section Review Hydrogen Ions And Acidity
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Section Review Hydrogen Ions And Acidity
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 19 2 Section Review Hydrogen Ions And Acidity
 - Setting Reading Goals 19 2 Section Review Hydrogen Ions And Acidity
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 19 2 Section Review Hydrogen Ions And Acidity

- Fact-Checking eBook Content of 19 2 Section Review Hydrogen Ions And Acidity
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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