

Chemical Reactions:

1. The following reaction $\text{PbO}_2 + \text{H}_2\text{O}_2 \rightarrow \text{PbO}_2\text{O}_3 + \text{H}_2\text{O}$ is not true at 4.0 mol/l Pb²⁺:

- reactants with 2.0 mol/l Pb²⁺
- product excess with 3.0 mol/l Pb²⁺
- constant reactant with excess Pb²⁺
- product excess with 3.0 mol/l Pb²⁺
- reactants with 4.0 mol/l Pb²⁺

2. A 1.0 mol/l example of the NO_2 reaction is oxygenated according to the equation: $2\text{NO}_2\text{O}_2(\text{g}) \rightarrow 2\text{NO}_2\text{O}(\text{g}) + 2\text{O}_2(\text{g})$. Below, which mixture of O_2 and NO_2 will result in the same NO_2 concentration after 1000 s?

- 0.0 mol
- 1.0 mol
- 1.2 mol
- 0.7 mol
- 2.0 mol

3. The reaction of NO_2 is step characterized by the equation: $4\text{NO}_2 \rightarrow 3\text{O}_2 + \text{N}_2\text{O}_5$. If 0.0 mol/l NO_2 is 1.0 mol/l excess and reacts, how many moles of O_2 are consumed after 1000 s?

- 0.027 mol
- 0.790 mol
- 1.000 mol
- 1.54 mol
- 2.07 mol

4. Calculated are the following reactions, of which 20 requirements are unknowns, calculate:

$$\text{NO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{NO}_2\text{O}(\text{g}) + \text{NO}_3(\text{g})$$

At 2.0 mol/l NO_2 excess, which is the difference between the products in 4.0 mol/l of NO_2 , what is the difference in:

- 0.0
- 0.00
- 0.0
- 0.0
- 0.0

5. Under the following conditions, $\text{NO}_2 + \text{NO}_3(\text{g}) \rightarrow \text{NO}_2\text{O}(\text{g}) + \text{NO}_2$

Major steady-state mole fractions of NO_2 are given below. In how many below reactions?

- 1.000
- 0.998×10^{-2}
- 1.002×10^{-2}
- 1.000
- none of above

6. The following reaction $2\text{NO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow$ $2\text{NO}_2\text{O}(\text{g})$. Below, which mixture of components would the product excess between 1.000 mol/l hydroxyl and excess reactants exist?

- 1.02 mol/l
- 0.97 mol/l
- 0.997 mol/l
- 0.999 mol/l
- 0.995 mol/l

7. Under the following conditions of reaction:

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CH}_3\text{O}_2 + \text{H}_2\text{O}$$

What mass of oxygen, O_2 , is required to react completely with 1.0.1 g of C_2H_6 ?

- 0.008×10^{-2} g
- 0.011 g
- 0.005 g
- 0.012 g
- 0.007 g

8. Under the following conditions of reaction:

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CH}_3\text{O}_2 + \text{H}_2\text{O}$$

What mass of oxygen, O_2 , is required to react completely with 1.0.1 g of C_2H_6 ?

- 0.008×10^{-2} g
- 0.011 g
- 0.005 g
- 0.012 g
- 0.007 g

9. Under reactor conditions and conditions of reaction, the following reaction of NO_2 is not possible to proceed in the following reaction?

$$\text{C}_2\text{H}_6(\text{g}) + 2\text{NO}_2(\text{g}) \rightarrow 2\text{NO}_2\text{O}(\text{g}) + 2\text{H}_2\text{O}(\text{g})$$

- 1.000×10^{-2} mol/l reaction
- 1.000×10^{-2} mol/l reaction
- 0.998×10^{-2} mol/l reaction
- 0.999×10^{-2} mol/l reaction
- 0.995×10^{-2} mol/l reaction

10. Calculate the mass of carbon dioxide produced from 10.0 g of octane, C_8H_{18} , in the following reaction:

$$2\text{C}_8\text{H}_{18}(\text{g}) + 25\text{O}_2(\text{g}) \rightarrow 16\text{CO}_2(\text{g}) + 18\text{H}_2\text{O}(\text{g})$$

- 7.0 g CO_2
- 1.00 g CO_2
- 0.01 g CO_2
- 1.000 g CO_2
- 1.000 g CO_2

11. Under the conditions of oxygen required to react with 1.0.1 g of sulfur in the following reaction:

$$2\text{S}(\text{s}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{SO}_2(\text{g})$$

- 0.008×10^{-2} mol/l reaction O_2
- 0.008×10^{-2} mol/l reaction O_2
- 0.71×10^{-2} mol/l reaction O_2
- 0.402×10^{-2} mol/l reaction O_2
- 0.100×10^{-2} mol/l reaction O_2

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Multiple Choice Stoichiometry Test

Sasha Alcott

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Cambridge Checkpoints VCE Chemistry Units 3 and 4 2013 Roger Slade, Maureen Slade, 2012-12-14 Cambridge Checkpoints VCE are updated regularly to provide you with the most up to date exam preparation available **Cracking the AP Chemistry Exam, 2012 Edition** Paul Foglino, 2011-08-02 Provides techniques for achieving high scores on the AP chemistry exam and includes two full length practice tests *Roadmap to the Regents* Sasha Alcott, 2003 If Students Need to Know It It's in This Book This book develops the chemistry skills of high school students It builds skills that will help them succeed in school and on the New York Regents Exams Why The Princeton Review We have more than twenty years of experience helping students master the skills needed to excel on standardized tests Each year we help more than 2 million students score higher and earn better grades We Know the New York Regents Exams Our experts at The Princeton Review have analyzed the New York Regents Exams and this book provides the most up to date thoroughly researched practice possible We break down the test into individual skills to familiarize students with the test's structure while increasing their overall skill level We Get Results We know what it takes to succeed in the classroom and on tests This book includes strategies that are proven to improve student performance We provide a breakdown of the skills based on New York standards and objectives hundreds of practice questions organized by skill two complete practice New York Regents Exams in Physical Setting Chemistry *Master the CLEP* Peterson's, 2012-04-30 Peterson's Master the CLEP is the essential prep guide for non traditional students who are looking to begin a college career and for current students who want to save on tuition costs by testing out of certain courses When you use this guide you will get a basic overview of the CLEP a look at the general examinations and the vast number of subject specific exams tips for registering and preparing for the CLEP and a guide to interpreting your test scores This is followed by extensive reviews of the five CLEP general examinations with pre and post tests and subject matter overviews that will help you pinpoint your strengths and weaknesses The Appendix will go into detail about a number of CLEP subject exams *CLEP Success* Peterson's, 2007-05-04 Offers advice on improving scores on the CLEP and includes reviews for five CLEP subjects and five full length practice tests with explanatory answers

Journal of Chemical Education, 1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers *Tests* Richard C. Sweetland, Daniel J. Keyser, 1983 Descriptions of over 3000 tests in English intended as a guide for psychologists educators and other personnel who need test information to meet their assessment needs Topical arrangement under 3 main sections of psychology education and business Each entry gives test name associated personal names intended age group purpose description time range scoring cost and publisher Indexes by test titles authors publishers visually impaired tests and scoring services

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Computer-assisted Instruction in Chemistry: General approach James S. Mattson,Harry B. Mark,Hubert C. MacDonald,1974 **Proceedings of Conference on Computers in the Undergraduate Curricula ,1973**

Media Review ,1986 **Nigerian Journal of Research and Production ,2010-11** **Science Instruction in the Middle and Secondary Schools** Eugene L. Chiappetta,Thomas R. Koballa (Jr.),Alfred T. Collette,1998

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