

# AP Chemistry Final Study Guide

## Unit 1: Atomic Structure and Properties (7%-9%)

- At STP, one mole of gas always occupies 22.4 L
- **Mass Spectroscopy:**
  - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
  - The most abundant isotope will have the greatest effect
- **Coulomb's Law:**
  - The greater the charge of the nucleus, the more energy is required to ionize an electron
  - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- **Electromagnetic Radiation:**
  - When atoms absorb energy, electrons jump to higher energy levels
  - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- **Kinetic Energy of Ionized Electrons:**
  - The faster an ejected electron is going, the more kinetic energy it has
  - Electrons originally further away from the nucleus require less energy to eject, so it will have more kinetic energy
- **Electron Configuration:**
  - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
    - Nickel:  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^8 4s^2$  (notice the d-block)
    - Nickel (with shorthand):  $[\text{Ar}]4s^2 3d^8$
- **Law of Definite Proportions:**
  - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- **Law of Multiple Proportions:**
  - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- **Aufbau Principle:**
  - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

# Ap Chem Study Guide

**Tpb Publishing**

## Ap Chem Study Guide:

ASAP Chemistry: a Quick-Review Study Guide for the AP Exam The Princeton Review,2018-11-15 Subtitle on cover A quick review study guide for the AP Exam **AP Chemistry** P. E. Dumas,R. M. Fikar,J. W. Samples,W. C. Uhland,2008-01-22 REA Real review Real practice Real results Get the college credits you deserve AP CHEMISTRY 10th Edition Completely aligned with today s AP exam Are you prepared to excel on the AP exam Set up a study schedule by following our results driven timeline Take the first practice test to discover what you know and what you should know Use REA s advice to ready yourself for proper study and success Practice for real Create the closest experience to test day conditions with 6 full length practice tests Chart your progress with full and detailed explanations of all answers Boost your confidence with test taking strategies and experienced advice Sharpen your knowledge and skills The book s full review features coverage of all AP Chemistry main topic areas such as solutions stoichiometry kinetics and thermodynamics as well as all subject areas found on the official exam including the structure of matter the states of matter chemical reactions and descriptive chemistry Smart and friendly lessons reinforce necessary skills Key tutorials enhance specific abilities needed on the test Targeted drills increase comprehension and help organize study Ideal for Classroom Family or Solo Test Preparation REA has provided advanced preparation for generations of advanced students who have excelled on important tests and in life REA s AP study guides are teacher recommended and written by experts who have mastered the course and the test

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