

GENERAL

- Organic molecules : contains C, forms backbone and branched structures of large diverse molecules as rings or chains, **MUST** form 4 covalent bonds
- Hydrocarbons : contains C & H & sometimes O, alkanes (C_nH_{2n+2}), alkenes (C_nH_{2n}), alkynes (C_nH_{2n-2})

1.1 THE CHEMISTRY OF LIFE

- Know what intramolecular and intermolecular forces are

Intermolecular Forces: attraction **BETWEEN** molecules (e.g. hydrogen bonding), like "outer"

Intramolecular Forces: attractions/forces within the molecules (e.g. ionic, polar covalent and covalent bonding)

- Know the difference between covalent and ionic

Ionic: e^- transfer; where resulting ions attract (metal \rightarrow non-metal)

Covalent: Atoms share e^- to fill valence shell for stable octet (generally 2 non-metals)

- Know what electronegativity is and what electronegativity difference is

Electronegativity: The measure of an atom's attraction to shared electrons

EN Difference: Determines what type of bond there is

- 0.0 – 0.4 : slightly/non-polar covalent (electronically symmetrical, X : X)
- 0.4 – 1.7 : polar covalent (partial charges, S^+ & S^-)
- 1.7 + : ionic bond (full charges, $M^+ X^-$)

- Know the difference between polar and nonpolar

Polarity - partial pos or neg charge @ ends of molecule (one with most EN has partial neg)

Polar: An unequal sharing of electrons; one atom attracts the electrons more than the other does

Non-polar: balanced out, usually gases

- **Hydrogen bond** – the attractive force between the partially charged H of one molecule & a partially negative charged atom in another molecule
- Water molecule has bent shape due to electrons on O pushing H's down 

- Know what an isotope and radioisotope is

Isotope: # of p^+ same, # of n^0 different, thus the mass # is different

Radioisotope: Radioactive version, decomposes over time

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