

## GENERAL

- Organic molecules : contains C, forms backbone and branched structures of large diverse molecules as rings or chains, **MUST** form 4 covalent bonds
  - Hydro-carbons : contains C & H & sometimes O, alkanes ( $C_nH_{2n+2}$ ), alkenes ( $C_nH_{2n}$ ), alkynes ( $C_nH_{2n-2}$ )
- 

## 1.1 THE CHEMISTRY OF LIFE ✓

- Know what intramolecular and intermolecular forces are

Intermolecular Forces: attraction **BETWEEN** molecules (e.g. hydrogen bonding), like "outer"

Intramolecular Forces: attractions/forces within the molecules (e.g. ionic, polar covalent and covalent bonding)

- Know the difference between covalent and ionic

Ionic:  $e^-$  transfer; where resulting ions attract (metal → non-metal)

Covalent: Atoms share  $e^-$  to fill valence shell for stable octet (generally 2 non-metals)

- Know what electronegativity is and what electronegativity difference is

Electronegativity: The measure of an atom's attraction to shared electrons

EN Difference: Determines what type of bond there is


- 0.0 — 0.4 : slightly/non-polar covalent (electronically symmetrical, X : X)
- 0.4 — 1.7 : polar covalent (partial charges,  $S^+ \& S^-$ )
- 1.7 + : ionic bond (full charges,  $M^+ X^-$ )

- Know the difference between polar and nonpolar

Polarity - partial pos or neg charge @ ends of molecule (one with most EN has partial neg)

Polar: An unequal sharing of electrons; one atom attracts the electrons more than the other does

Non-polar: balanced out, usually gases

- **Hydrogen bond** - the attractive force between the partially charged H of one molecule & a partially negative charged atom in another molecule
- Water molecule has bent shape due to electrons on O pushing H's down 

- Know what an isotope and radioisotope is

Isotope: # of  $p^+$  same, # of  $n^0$  different, thus the mass # is different

Radioisotope: Radioactive version, decomposes over time

# Bio Study Guide Biochemistry

**United States. Court of Claims**



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## **Table of Contents Bio Study Guide Biochemistry**

1. Understanding the eBook Bio Study Guide Biochemistry
  - The Rise of Digital Reading Bio Study Guide Biochemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Bio Study Guide Biochemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Bio Study Guide Biochemistry
  - User-Friendly Interface

4. Exploring eBook Recommendations from Bio Study Guide Biochemistry
  - Personalized Recommendations
  - Bio Study Guide Biochemistry User Reviews and Ratings
  - Bio Study Guide Biochemistry and Bestseller Lists
5. Accessing Bio Study Guide Biochemistry Free and Paid eBooks
  - Bio Study Guide Biochemistry Public Domain eBooks
  - Bio Study Guide Biochemistry eBook Subscription Services
  - Bio Study Guide Biochemistry Budget-Friendly Options
6. Navigating Bio Study Guide Biochemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Bio Study Guide Biochemistry Compatibility with Devices
  - Bio Study Guide Biochemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Bio Study Guide Biochemistry
  - Highlighting and Note-Taking Bio Study Guide Biochemistry
  - Interactive Elements Bio Study Guide Biochemistry
8. Staying Engaged with Bio Study Guide Biochemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Bio Study Guide Biochemistry
9. Balancing eBooks and Physical Books Bio Study Guide Biochemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Bio Study Guide Biochemistry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Bio Study Guide Biochemistry
  - Setting Reading Goals Bio Study Guide Biochemistry
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Bio Study Guide Biochemistry
  - Fact-Checking eBook Content of Bio Study Guide Biochemistry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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