

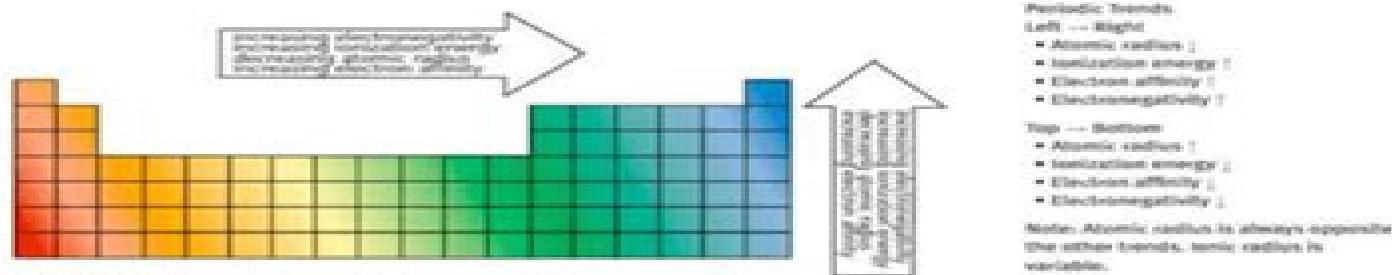
Chemistry:

CHAPTER 1

- Charge of electron $\approx 1.6 \times 10^{-19}$; Avogadro's # $\approx 6 \times 10^{23}$
- Mass number (A) on top of element, atomic number (Z) below
- $E = \frac{hc}{\lambda} = -R \left[\frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$; $E = qV$ (energy = charge on e- x electron volts); $c = \lambda f$ ($c = 3 \times 10^8$)
- Heisenberg uncertainty — impossible to determine momentum and position simultaneously
- Pauli exclusion — no electrons have same quantum #’s
- Max e- in a shell $\approx 2n^2$, max e- in subshell $\approx 4l+2$
 - Possible values for l ≈ 0 to $n-1$, possible values for ml $\approx -l$ to $+l$
- S can hold 2 e-, p ≈ 6 , d ≈ 10 , f ≈ 14
- Aufbau principle (building up) — lowest n+l will fill first
- For anions add extra e-, for cations remove e- from subshells with highest n first
- Hund’s — half filled orbitals with parallel spin e-, attracted to magnetic field
 - Diamagnetic — paired e-, repelled by magnetic field
 - Ferromagnetic — unpaired e-, randomly oriented dipoles, become magnetized in field
- A faraday is equal to one mole of electric charge

CHAPTER 2

- Z_{eff} increases from left to right, constant up and down
- Ionic radius is smallest near metalloid line and increases out
- F is smallest, most EN, highest IE and EA; Cs is the largest and least EN
- Complex ions from transition metals associate with water to form **hydration complexes**
- Atomic radius has the opposite trend of all others



CHAPTER 3

- C, N, O, F, Na, Mg always abide by octet rule; H, He, Li, Be, B can have incomplete; period 3 or higher can have expanded
- **Coordinate covalent bond** — both e- are contributed by one atom (lone pair attack, acid-base)
- Ionic ΔEN difference greater than 1.7, polar covalent ≈ 0.5 to 1.7, nonpolar ΔEN less than 0.5
 - Ionic are stronger than covalent (which are both stronger than IM forces)
 - Van der Waals occurs more with close proximity and large molecules
- **Dipole moment** $\approx p = qd$
- Triple bonds are shorter and have more energy
- Major resonance contributor is more stable (less formal charge, negative on more EN atom)
 - Formal charge = valence - nonbonding - $\frac{1}{2}$ bonding
- Electron geometry is when bonding and lone are equal, molecular geo is when lone repel more

Section 2 nuclear Chemistry Study Guide Key

Susan McMurry

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Table of Contents Section 2nuclear Chemistry Study Guide Key

1. Understanding the eBook Section 2nuclear Chemistry Study Guide Key
 - The Rise of Digital Reading Section 2nuclear Chemistry Study Guide Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 2nuclear Chemistry Study Guide Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 2nuclear Chemistry Study Guide Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 2nuclear Chemistry Study Guide Key
 - Personalized Recommendations
 - Section 2nuclear Chemistry Study Guide Key User Reviews and Ratings
 - Section 2nuclear Chemistry Study Guide Key and Bestseller Lists
5. Accessing Section 2nuclear Chemistry Study Guide Key Free and Paid eBooks
 - Section 2nuclear Chemistry Study Guide Key Public Domain eBooks
 - Section 2nuclear Chemistry Study Guide Key eBook Subscription Services
 - Section 2nuclear Chemistry Study Guide Key Budget-Friendly Options
6. Navigating Section 2nuclear Chemistry Study Guide Key eBook Formats
 - ePUB, PDF, MOBI, and More
 - Section 2nuclear Chemistry Study Guide Key Compatibility with Devices
 - Section 2nuclear Chemistry Study Guide Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 2nuclear Chemistry Study Guide Key
 - Highlighting and Note-Taking Section 2nuclear Chemistry Study Guide Key
 - Interactive Elements Section 2nuclear Chemistry Study Guide Key
8. Staying Engaged with Section 2nuclear Chemistry Study Guide Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Section 2nuclear Chemistry Study Guide Key

9. Balancing eBooks and Physical Books Section 2nuclear Chemistry Study Guide Key

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Section 2nuclear Chemistry Study Guide Key

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Section 2nuclear Chemistry Study Guide Key

- Setting Reading Goals Section 2nuclear Chemistry Study Guide Key
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Section 2nuclear Chemistry Study Guide Key

- Fact-Checking eBook Content of Section 2nuclear Chemistry Study Guide Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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