

Chemistry 2nd Semester Exam Review Sheet:

Chapters 9, 10, 11, 12, 13, 14, 15, and 16

Chapter 9: Chemical Quantities

1. Solve stoichiometry problems using grams \rightarrow moles \rightarrow moles \rightarrow grams calculations. Examples below:



$$30.0 \text{ moles} \times 1 \text{ mol}/2 \text{ mol} = 15.0 \text{ moles}$$

$$15.0 \text{ moles } (\text{NH}_4)_2\text{SO}_4 \times 132.2 \text{ gram/mole} = 1983 = 1980 \text{ grams (to 3 s.f.)}$$

- b. $\text{Na}_2\text{CO}_3 + 2\text{HNO}_3 \rightarrow 2\text{NaNO}_3 + \text{CO}_2 + \text{H}_2\text{O}$. How many moles/grams of sodium nitrate are produced by the complete reaction of 100. grams of sodium carbonate? Moles of nitric acid needed to produce 500.0 grams of sodium nitrate? *Be able to calculate grams or moles of any reactant or product given a quantity of reactant or product.*

Grams \rightarrow Moles \rightarrow Moles \rightarrow Grams

$$100. \text{ g Na}_2\text{CO}_3 \times 1 \text{ mol}/106.0 \text{ g} \times 2 \text{ mol NaNO}_3 / 1 \text{ mol Na}_2\text{CO}_3 \times 85.00 \text{ g/mol} = 160.4 \text{ g}$$

$$500.0 \text{ g NaNO}_3 \times 1 \text{ mol}/85.00 \text{ g} \times 2 \text{ mol}/2 \text{ mol} \times 63.02 \text{ g/mol} = 370.7 \text{ grams}$$

2. Calculate percent yield of products. Example below using (b) above.
What is the theoretical yield of NaNO_3 if 100.0 grams of Na_2CO_3 are used in the reaction? What is the percent yield of NaNO_3 if 130.5 grams of NaNO_3 are actually collected?

Theoretical yield = 160.4 grams (answer from above)

$$\text{Percent Yield} = (\text{Actual Yield/Theoretical Yield}) \times 100 = (130.5 \text{ g}/160.4 \text{ g}) \times 100 = 81.36\%$$

3. What is a limiting reactant in a chemical reaction? The reactant that is used up first, thereby limiting the amount of product that can be formed.

Chapter 10: Modern Atomic Theory

4. Describe Rutherford's model of the atom. What could it not explain?

Nucleus (very small) consisting of protons with electrons in orbits around the nucleus. Could not explain how the electrons remained in orbit.

5. Define electromagnetic radiation, wavelength, and frequency.

6. How do atoms gain energy to become excited? **Absorb a photon of energy**

How do they lose energy to return to their ground state? **Release a photon of same energy as that absorbed**

7. What is the evidence that energy absorbed by the hydrogen atom is quantized? **Specific wavelengths of light seen in emission spectra for atoms. (page 285-86)**

8. Describe the Bohr model of the atom. How did it improve on Rutherford's model?

Same as Rutherford with respect to nucleus. Electrons located on orbits at specific distances from nucleus.

9. Draw the orbital diagram (with the arrows) for an element: for example, phosphorus, sodium, and fluorine.

See book, notes, probs 54, 95

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Official Gazette Philippines, 2004 **The Universal tutor** London gen. corresp. coll, 1894 *Directory of Distance Learning Opportunities* Modoc Press, Inc., 2003-02-28 This book provides an overview of current K 12 courses and programs offered in the United States as correspondence study or via such electronic delivery systems as satellite cable or the Internet The Directory includes over 6 000 courses offered by 154 institutions or distance learning consortium members Following an introduction that describes existing practices and delivery methods the Directory offers three indexes Subject Index of Courses Offered by Level Course Level Index Geographic Index All information was supplied by the institutions Entries include current contact information a description of the institution and the courses offered grade level and admission information tuition and fee information enrollment periods delivery information equipment requirements credit and grading information library services and accreditation *The Journal of Education* , 1902 Oxford University Gazette University of Oxford, 1910 **Guy's Hospital Gazette** , 1904 **Journal of Gas Lighting and Water Supply** , 1926 *Federation Bulletin* , 1918 Tests in Print Oscar Krisen Buros, 1983 *The Medical Directory* , 1903 **Medical Education, Medical Colleges and the Regulation of the Practice of Medicine in the United States and Canada** , 1891 British Medical Journal , 1929 An Experimental Study of the Iowa Placement Examinations Lawrence Williams Miller, University of Iowa, 1930 **The ETS Test Collection Catalog** Educational Testing Service. Test Collection, 1993 The major source of information on the availability of standardized tests Wilson Library Bulletin Covers commercially available standardized tests and hard to locate research instruments **Announcements for the Year ...** Purdue University, 1920 *The Annual Catalogue of Purdue University, Lafayette, Indiana ... with Announcements for ...* Purdue University, 1922 **The Educational year book. [5 issues].** , 1882 Accounting and Auditing Examinations Arco Publishing Company, 1949 Chemical News and Journal of Industrial Science , 1901 *Glasgow University Calendar* University of Glasgow, 1914

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