



# HYDROGEN IONS AND ACIDITY

## Section Summary

### Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

### Key Vocabulary

<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li></ul>	<ul style="list-style-type: none"><li>concentration</li><li>hydrogen ion</li><li>ion</li><li>logarithm</li><li>logarithmic scale</li><li>acid</li><li>strong acid</li><li>weak acid</li></ul>
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### Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

### Key Figures

- Figure 1: pH scale
- Figure 2: Ionization of water

### Principles of Ionization

Water is a neutral molecule in which the hydrogen and oxygen atoms are covalently bonded together. It is the ionization of these atoms that gives water its acidic properties.

Water ionizes to a very small extent to form a very small concentration of hydrogen ions ( $\text{H}^+$ ) and hydroxide ions ( $\text{OH}^-$ ). The concentration of hydrogen ions is  $10^{-7}$  M.

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### Ionization

<ul style="list-style-type: none"><li>Water ionization</li><li>Acid</li><li>Strong acid</li><li>Weak acid</li><li>Neutral</li><li>Basic</li><li>Strong base</li><li>Weak base</li></ul>	<ul style="list-style-type: none"><li>Water ionization</li><li>Acid</li><li>Strong acid</li><li>Weak acid</li><li>Neutral</li><li>Basic</li><li>Strong base</li><li>Weak base</li></ul>
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