

CHAPTER 7**STUDY GUIDE****Ionic Compounds and Metals****Section 7.1 Ion Formation**

In your textbook, read about chemical bonds and formation of ions.

Use each of the terms below just once to complete the passage.

chemical bond	electrons	energy level	ions	noble gases
nucleus	octet	pseudo-noble gas formations	valence	

The force that holds two atoms together is called a(n) (1) _____.

Such an attachment may form by the attraction of the positively charged

(2) _____ of one atom for the negatively charged

(3) _____ of another atom, or by the attraction of charged atoms, which are called

(4) _____. The attractions may also involve

(5) _____ electrons, which are the electrons in the outermost

(6) _____. The (7) _____ are a family of elements that have very

little tendency to react. Most of these elements have a set of eight outermost electrons, which is called a

stable (8) _____. The relatively stable electron structures developed by loss of

electrons in certain elements of groups 3, 4, 13, and 14 are called (9) _____.

For each statement below, write *true* or *false*.

10. A positively charged ion is called an anion.
11. Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge.
12. Elements tend to react so that they acquire the electron structure of a halogen.
13. A sodium atom tends to lose one electron when it reacts.
14. The electron structure of a zinc ion (Zn^{2+}) is an example of a pseudo-noble gas formation.
15. A Cl^- ion is an example of a cation.
16. The ending *-ide* is used to designate an anion.
17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions.

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Chemical Principles Student's Study Guide & Solutions Manual John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better. The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text. **Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon buckyballs, graphite, and diamond are illustrated at the left as is the molecule methane CH₄ from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement.

Science Mastermind Class 10 : A Comprehensive Study Guide with Diagrams and Practice Exercise's Gaurav Sinha, 2025-04-04 This book is helpful for class 10 cbse science students.

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Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text. This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises. New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application.

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