

## Student Exploration: Calorimetry Lab

(Note: Gizmos can be done even on your phones! Contact Goff or Molik if you need your log-in information)

### Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance's **specific heat capacity**. You will use the *Calorimetry Lab* Gizmo to determine the specific heat capacities of various substances.



- On the **SIMULATION** pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the **GRAPH** tab, and click **Play** (▶).

- What was the **Final temperature** of the copper and the water? 35 °C \_\_\_\_\_
- How much did the temperature of the copper change? It went down 65 °C \_\_\_\_\_
- How much did the temperature of the water change? \_\_\_\_\_ it went up 5 °C \_\_\_\_\_

- Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

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<b>Activity A:</b> Heat transfer	<u>Get the Gizmo ready:</u> • Click <b>Reset</b> (↺).	
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**Question:** What factors determine how heat energy transfers between objects?

- Predict: In the Gizmo warm-up, you saw how 200 g of 90 °C copper transfers heat to 200 g of 30.0 °C water.

- How do you think increasing the water's mass would affect the final temperature?

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# Answer Key Explore Learning Calorimetry Lab

**David Kirk**



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