

# ALGEBRA I END OF COURSE STUDY GUIDE

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**Slope**  $y = mx + b$   
*m is the slope*  
*b is the y-axis intercept*

Find Slope given a table:



Find Slope given a graph:



Linear Functions:

Possible representations

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Quadratics: Parabolic function  $y = ax^2 + bx + c$



Exponent Basics:  $x^a \cdot x^b = x^{a+b}$ ,  $x^a \div x^b = x^{a-b}$ ,  $x^a = x^{a/1}$

Add Exponents & multiply numbers

$$x \cdot x = x^2$$

$$x^2 \cdot x^2 = x^4$$

$$x^3 \cdot x^2 = x^5$$

$$x^4 \cdot x^3 = x^7$$

$$x^5 \cdot x^2 = x^7$$

$$x^6 \cdot x^1 = x^7$$

$$x^7 \cdot x^0 = x^7$$

Subtract Exponents & Divide numbers

$$\frac{x^5}{x^2} = x^{5-2} = x^3$$

$$\frac{27x^3y^2}{3x^2y^2} = \frac{27}{3} \cdot \frac{x^3}{x^2} \cdot \frac{y^2}{y^2} = 9x^{3-2} \cdot 1 = 9x$$

$$\frac{(-6x^2y^3)(2x^2y^3)}{-18x^2y^3} = \frac{-12x^4y^6}{-18x^2y^3} = \frac{2}{3}x^{4-2}y^{6-3} = \frac{2}{3}x^2y^3$$

Distributive property

$$4(x + 3) = 4x + 12$$

$$-3x(2x - 4) + 6x = (-3x)(2x) + (-3x)(-4) + 6x$$

$$= -6x^2 + 12x + 6x = -6x^2 + 18x$$

$$-2(3x^2 + 5x) + 12x = -2(3x^2) + -2(5x) + 12x$$

$$= -6x^2 - 10x + 12x = -6x^2 + 2x$$

$$\frac{2}{5}(15ab^2 - 20ab^2) - \frac{2}{3}(32ab^2 - 6ab^2)$$

$$= \frac{2}{5}(15ab^2 - 20ab^2) - \frac{2}{3}(32ab^2 - 6ab^2)$$

$$= 6ab^2 - 8ab^2 - \frac{64}{3}ab^2 + 4ab^2$$

$$= -2ab^2 - \frac{64}{3}ab^2 + 4ab^2$$

$$= -2ab^2 - \frac{64}{3}ab^2 + \frac{12}{3}ab^2$$

$$= -2ab^2 - \frac{52}{3}ab^2$$

$$= -\frac{58}{3}ab^2$$

Combine Like Terms  
 Same Variables with the Same Exponents

Find the slope of the line passing through  $(-1, 2)$  and  $(3, 4)$

$m = \frac{y_2 - y_1}{x_2 - x_1} = \frac{4 - 2}{3 - (-1)} = \frac{2}{4} = \frac{1}{2}$

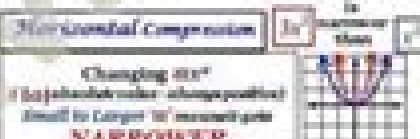
Equation of the line:  $y - 2 = \frac{1}{2}(x + 1)$

$y - 2 = \frac{1}{2}x + \frac{1}{2}$

$y = \frac{1}{2}x + \frac{5}{2}$



Transformations of Quadratic Functions



# Algebra 1 Study Guide With Answers

**Miriam A. Leiva**



## **Algebra 1 Study Guide With Answers:**

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**Algebra 1: Explorations and Applications** Holt McDougal,2001-01      El-Hi Textbooks and Serials in Print ,2006

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monoxide co determine the mass of iodine  $I_2$  which could be produced b if in the above situation only 0.160 moles of iodine  $I_2$  was produced i what mass of iodine was produced

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $NH_3$  can be produced theoretically from the reaction of 5.0 g of n what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation  $C_3H_6 + 3O_2 \rightarrow 3CO + 3H_2O$  if the reaction of 21.3 grams of  $O_2$  produces a 47.3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent

you have 0.216232 grams mole 21.3 grams 0.66 moles 32 g m

experiment 4 stoichiometry limiting reagents yield making - Feb 26 2022

web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained

calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

w limiting reagents and percent yield hw1 answerkey - Nov 06 2022

web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  4  $\text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0

g  $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$  4 mol  $\text{Al} \times 1 \text{ mol Al}_2\text{O}_3$  c determine the percent yield

yield actual  $\times 100$   $26.0 \text{ g} \times 100 / 86.01 \text{ theoretical} = 302.3 \text{ g}$

**7 2 limiting reagent and reaction yields chem 1114** - Jul 02 2022

web 7  $\text{Na}_2\text{CO}_3$  is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be

weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

C is the limiting reagent 4 33 g of  $\text{H}_2$  are left over 13

6 5 limiting reagent and percent yield chemistry libretexts - Aug 15 2023

web learning objectives define and determine theoretical yields actual yields and percent yields identify a limiting reagent

from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s

remains when the reaction is complete

**limiting reactant and reaction yields article khan academy** - Jan 08 2023

web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent

yield theoretical  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  based on this definition we would expect a percent yield

**8 5 limiting reactant theoretical yield and percent yield** - Dec 07 2022

web may 20 2018 learning objectives identify the limiting reactant limiting reagent in a given chemical reaction calculate

how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is

complete

**12 3 limiting reagent and percent yield worksheet answer key** - Jan 28 2022

web limiting reagent and percent yield practice worksheet answer key limiting reactant yield practice worksheet 1 methanol

$\text{CH}_3\text{OH}$  can be produced by the following reaction  $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$  calculate the theoretical yield of  $\text{CH}_3\text{OH}$  if 68.5 g of CO is

reacted with 8.6 g of  $\text{H}_2$

limiting reactant and percent yield worksheet with key - Oct 05 2022

web honors chemistry 1b limit reactant and percent yield worksheet with excess calculation name sto 4 solve stoichiometric

problems from a balanced chemical equation 3 sto 5 differentiate between the actual yield and theoretical yield of a chemical

reaction

*limiting reactant and percent yield practice hs science* - Jun 01 2022

web limiting reactant and percent yield practice name 1 consider the following reaction  $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{PO}_4 + \text{NaNO}_3$  which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate what is the

*limiting reagents and percentage yield worksheet answers* - Sep 04 2022

web limiting reagents and percentage yield worksheet 1 consider the reaction  $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow \text{I}_2 + 5\text{CO}_2$  a 80.0 grams of iodine v oxide  $\text{I}_2\text{O}_5$  reacts with 28.0 grams of carbon monoxide  $\text{CO}$  determine the mass of iodine  $\text{I}_2$  which could be produced 80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$   $\times$  1 333.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

**7.3 limiting reactant and percent yield problems** - Jun 13 2023

web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3