

Chemistry Study Guide

Periodic Table of the Elements

<small>1</small> H <small>1.008</small>																	<small>2</small> He <small>4.003</small>																															
<small>3</small> Li <small>6.941</small>	<small>4</small> Be <small>9.012</small>																	<small>5</small> B <small>10.81</small>	<small>6</small> C <small>12.01</small>	<small>7</small> N <small>14.01</small>	<small>8</small> O <small>16.00</small>	<small>9</small> F <small>18.998</small>	<small>10</small> Ne <small>20.18</small>																									
<small>11</small> Na <small>22.99</small>	<small>12</small> Mg <small>24.31</small>																	<small>13</small> Al <small>26.98</small>	<small>14</small> Si <small>28.09</small>	<small>15</small> P <small>30.97</small>	<small>16</small> S <small>32.07</small>	<small>17</small> Cl <small>35.45</small>	<small>18</small> Ar <small>39.95</small>																									
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Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles 🎯

Matter:

Anything that has **mass** and **occupies space** ✖️🌐

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** 🧪👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

Section 2nuclear Chemistry Study Guide Key

Darrell D. Ebbing



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Study Guide with Answers to Selected Problems H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

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