

# Chem 8 Practice

1. The balanced equation  $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$  tells us that 4.0 mol  $\text{H}_2$

- A) reacts with 2.0 mol  $\text{P}_4$ .
- B) produces 8.0 mol  $\text{PH}_3$ .
- C) cannot react with phosphorus.
- D) produces 2.7 mol  $\text{PH}_3$ .
- E) reacts with 4.0 mol  $\text{P}_4$ .

2. A 1.0-mol sample of  $\text{PCl}_5$  was decomposed according to the equation



How many moles of  $\text{Cl}_2$  are formed assuming 100% yield?

- A) 0.5 mol
- B) 1.0 mol
- C) 1.2 mol
- D) 0.7 mol
- E) 2.0 mol

3. The rusting of iron is represented by the equation  $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$ . If you have a 1.04-mol sample of iron, how many moles of  $\text{Fe}_2\text{O}_3$  will there be when the iron has rusted completely?

- A) 0.527 mol
- B) 0.790 mol
- C) 1.05 mol
- D) 1.56 mol
- E) 2.07 mol

4. Consider the following reaction, where X represents an unknown element.

$$5\text{X}_2(\text{s}) + 2\text{H}_2\text{O}_2(\text{g}) \rightarrow \text{H}_2\text{X}_2(\text{l}) + 2\text{XO}_2(\text{g})$$

If 175 g of X react completely with hydrogen peroxide to produce 1.43 mol of  $\text{H}_2\text{X}_2$ , what is the identity of X?

- A) Ge
- B) Mg
- C) Se
- D) Sn
- E) C

5. Refer to the following equation:



How many molecules of  $\text{NO}_2$  are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08
- B)  $3.64 \times 10^{23}$
- C)  $1.82 \times 10^{23}$
- D) 128
- E) none of these

6. In the reaction  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ , how many moles of ammonia would be produced from 1.08 mol of hydrogen and excess nitrogen?

- A) 1.42 mol
- B) 2.27 mol
- C) 0.727 mol
- D) 2.18 mol
- E) 0.543 mol

7. Refer to the following unbalanced equation:

$$\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

What mass of oxygen ( $\text{O}_2$ ) is required to react completely with 13.1 g of  $\text{C}_4\text{H}_{10}$ ?

- A)  $2.98 \times 10^2$  g
- B) 23.1 g
- C) 4.86 g
- D) 49.2 g
- E) 0.152 g

8. How many molecules of carbon dioxide will be formed if 7.07 g of propane is burned in the following reaction?



- A)  $1.37 \times 10^{23}$  molecules
- B)  $1.03 \times 10^{23}$  molecules
- C)  $4.14 \times 10^{23}$  molecules
- D)  $3.10 \times 10^{23}$  molecules
- E)  $5.17 \times 10^{23}$  molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane,  $\text{C}_8\text{H}_{18}$ , in the following reaction.



- A) 79.2 g  $\text{CO}_2$
- B) 15.8 g  $\text{CO}_2$
- C) 99.1 g  $\text{CO}_2$
- D) 1.238 g  $\text{CO}_2$
- E) 1.400 g  $\text{CO}_2$

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction.



- A)  $1.03 \times 10^{23}$  molecules  $\text{O}_2$
- B)  $2.28 \times 10^{23}$  molecules  $\text{O}_2$
- C)  $1.71 \times 10^{23}$  molecules  $\text{O}_2$
- D)  $3.42 \times 10^{23}$  molecules  $\text{O}_2$
- E)  $5.13 \times 10^{23}$  molecules  $\text{O}_2$

# Multiple Choice Stoichiometry Test

**Paul Foglino, Princeton Review (Firm)**

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the text box.

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*Cambridge Checkpoints VCE Chemistry Units 3 and 4 2013* Roger Slade, Maureen Slade, 2012-12-14 Cambridge Checkpoints VCE are updated regularly to provide you with the most up to date exam preparation available **Cracking the AP Chemistry Exam, 2012 Edition** Paul Foglino, 2011-08-02 Provides techniques for achieving high scores on the AP chemistry exam and includes two full length practice tests *Roadmap to the Regents* Sasha Alcott, 2003 If Students Need to Know It It's in This Book This book develops the chemistry skills of high school students It builds skills that will help them succeed in school and on the New York Regents Exams Why The Princeton Review We have more than twenty years of experience helping students master the skills needed to excel on standardized tests Each year we help more than 2 million students score higher and earn better grades We Know the New York Regents Exams Our experts at The Princeton Review have analyzed the New York Regents Exams and this book provides the most up to date thoroughly researched practice possible We break down the test into individual skills to familiarize students with the test's structure while increasing their overall skill level We Get Results We know what it takes to succeed in the classroom and on tests This book includes strategies that are proven to improve student performance We provide a breakdown of the skills based on New York standards and objectives hundreds of practice questions organized by skill two complete practice New York Regents Exams in Physical Setting Chemistry *Master the CLEP* Peterson's, 2012-04-30 Peterson's Master the CLEP is the essential prep guide for non traditional students who are looking to begin a college career and for current students who want to save on tuition costs by testing out of certain courses When you use this guide you will get a basic overview of the CLEP a look at the general examinations and the vast number of subject specific exams tips for registering and preparing for the CLEP and a guide to interpreting your test scores This is followed by extensive reviews of the five CLEP general examinations with pre and post tests and subject matter overviews that will help you pinpoint your strengths and weaknesses The Appendix will go into detail about a number of CLEP subject exams *CLEP Success* Peterson's, 2007-05-04 Offers advice on improving scores on the CLEP and includes reviews for five CLEP subjects and five full length practice tests with explanatory answers

**Journal of Chemical Education**, 1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers Tests Richard C. Sweetland, Daniel J. Keyser, 1983 Descriptions of over 3000 tests in English intended as a guide for psychologists educators and other personnel who need test information to meet their assessment needs Topical arrangement under 3 main sections of psychology education and business Each entry gives test name associated personal names intended age group purpose description time range scoring cost and publisher Indexes by test titles authors publishers visually impaired tests and scoring services

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**Resources in Education** ,1984 **McGraw-Hill's 15 Practice SAT Subject Tests** McGraw-Hill,2006-04-03 If You Want a Faster more efficient preparation for the SAT Subject Tests b Practice for tests on more than one subject c Strategies from the test prep professionals d Entrance into the best college possible e All of the above Then you need McGraw Hill s Practice SAT Subject Tests If you need to practice for more than one SAT Subject Test or if you just want to try a few samples to help decide which test to take McGraw Hill s 15 Practice SAT Subject Tests prepares you for toplevel performance It provides two practice exams for each of the five leading enrollment tests U S History Math Level 1 Math Level 2 Biology E M and Chemistry plus five additional SAT Subject Test samples in World History Physics English Literature Spanish and French Unique features to suit every student s needs include 15 sample tests on the most popular subjects Specific question answering strategies for the most common question types Invaluable information on the academic background you need for each test Packed with proven tips from test prep professionals McGraw Hill s SAT Subject Tests is the smartest way to build test taking confidence get higher scores and win admission to the college of your choice

**A Summary of Science Education, 1988** Leopold E. Klopfer,1990-06 *Three Cognitive Skills in Chemistry and Their Application to Stoichiometry* Ardas Ozsogomonyan,1977 Cracking the AP Chemistry, 2002-2003 Edition Paul Foglino,Princeton Review (Firm),2002-01-29 The Princeton Review realizes that acing the AP Chemistry exam is very different from getting straight As in school The Princeton Review doesn t try to teach students everything there is to know about chemistry only the techniques they ll need to score higher on the exam There s a big difference In Cracking the AP Chemistry TPR will teach test takers how to think like the test makers and Score higher by reviewing key chemistry concepts Earn more points by becoming familiar with the format of the test Safeguard against traps that can lower scores Perfect skills with review questions in each chapter This book includes 2 full length simulated AP Chemistry exams All of The Princeton Review practice test questions are like the ones test takers will see on the actual exam and every solution is fully explained

**Computer-assisted Instruction in Chemistry: General approach** James S. Mattson,Harry B. Mark,Hubert C. MacDonald,1974 Proceedings of Conference on Computers in the Undergraduate Curricula ,1973 Media Review ,1986 **Nigerian Journal of Research and Production** ,2010-11 **Science Instruction in the Middle and Secondary Schools** Eugene L. Chiappetta,Thomas R. Koballa (Jr.),Alfred T. Collette,1998

New edition of a text providing far more than simply a cookbook of activities for science teachers Coverage includes discussion on the nature of science national standards and innovative programs the nature of adolescent learners and their schools teaching strategies and classroom management pl *The Education Index* ,1984

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