

Chem 8 Practice

1. The balanced equation $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$ tells us that 4.0 mol H_2

- A) reacts with 2.0 mol P_4 .
- B) produces 8.0 mol PH_3 .
- C) cannot react with phosphorus.
- D) produces 2.7 mol PH_3 .
- E) reacts with 4.0 mol P_4 .

2. A 1.0-mol sample of PCl_5 was decomposed according to the equation



How many moles of Cl_2 are formed assuming 100% yield?

- A) 0.5 mol
- B) 1.5 mol
- C) 1.0 mol
- D) 0.7 mol
- E) 2.0 mol

3. The rusting of iron is represented by the equation $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$. If you have a 1.04-mol sample of iron, how many moles of Fe_2O_3 will there be when the iron has rusted completely?

- A) 0.527 mol
- B) 0.790 mol
- C) 1.05 mol
- D) 1.58 mol
- E) 2.07 mol

4. Consider the following reaction, where X represents an unknown element.

$$5\text{X}_2(\text{s}) + 2\text{H}_2\text{O}_2(\text{l}) \rightarrow \text{H}_2\text{X}_2(\text{l}) + 2\text{XO}_2(\text{g})$$

If 175 g of X react completely with hydrogen peroxide to produce 1.43 mol of H_2X_2 , what is the identity of X?

- A) Ge
- B) Mg
- C) Se
- D) Sn
- E) C

5. Refer to the following equation:



How many molecules of NO_2 are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08
- B) 3.64×10^{23}
- C) 1.82×10^{23}
- D) 128
- E) none of these

6. In the reaction $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$, how many moles of ammonia would be produced from 1.08 mol of hydrogen and excess nitrogen?

- A) 1.42 mol
- B) 3.27 mol
- C) 0.727 mol
- D) 2.18 mol
- E) 0.543 mol

7. Refer to the following unbalanced equation:

$$\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

What mass of oxygen (O_2) is required to react completely with 13.1 g of C_4H_{10} ?

- A) 2.98×10^2 g
- B) 23.1 g
- C) 4.86 g
- D) 49.2 g
- E) 0.152 g

8. How many molecules of carbon dioxide will be formed if 7.07 g of propane is burned in the following reaction?



- A) 1.37×10^{23} molecules
- B) 1.03×10^{23} molecules
- C) 4.14×10^{23} molecules
- D) 3.10×10^{23} molecules
- E) 5.17×10^{23} molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane, C_8H_{18} , in the following reaction.



- A) 79.2 g CO_2
- B) 158 g CO_2
- C) 99.1 g CO_2
- D) 1.238 g CO_2
- E) 1.400 g CO_2

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction.



- A) 1.03×10^{23} molecules O_2
- B) 2.28×10^{23} molecules O_2
- C) 1.71×10^{23} molecules O_2
- D) 3.42×10^{23} molecules O_2
- E) 5.13×10^{23} molecules O_2

Multiple Choice Stoichiometry Test

Paul Foglino, Princeton Review (Firm)

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the text box.

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web example 3 company merger letter company name or letterhead address city state zip date addressee address city state

zip dear fellow employees we are in the process of making a number of changes in our worldwide manufacturing operations that will require us to establish an after tax profit reserve of about 555 000 000

merger announcement letter 4 templates writolay - May 12 2023

web oct 7 2020 with this letter template you can happily announce the merger of your company with another company you can provide information about the organization that you are merging with you need to mention your organization details as well you can reveal how the merger is going to benefit the employees of the company

example letters for changing the ownership merger acquisition - Apr 30 2022

web example 1 letter for change of ownership merger company name or letterhead address city state zip date addressee address city state zip dear mr johnson as i am sure you are aware from recent news reports first bank and dayton credit union are joining hands to better serve you