

$$\text{Mol} = \text{mass} / \text{Mr}$$

$$\text{Mol} = \text{concentration (mol/dm}^3\text{)} \times \text{volume (dm}^3\text{)}$$

$$\text{Volume (dm}^3\text{)} = \text{mol} / 24$$

$$1 \text{ cm}^3 = 0.001 \text{ dm}^3 \text{ (divided by 1.000)}$$

What is a mol? One mole of a substance is equal to 6.022×10^{23} units of that substance (such as atoms, molecules, or ions). The number 6.022×10^{23} is known as Avogadro's constant.

$$\text{Avogadro constant} = 6.022 \times 10 \text{ to the power of } 23$$

Percentage mass:

Percentage mass = the percentage by mass of an element in the compound

$$\text{E.g \% by mass of Fe in Fe}_2\text{O}_3 = 100 \times 2(56) / 160 = 70\%$$

Moles:

E.g what mass of oxygen is reacted when 6g of magnesium reacts?



Step 1: Use the equation $\text{mol} = \text{mass} / \text{Mr}$ to find the number of moles of magnesium that reacted

$$6\text{g} / 24\text{g} = 0.25 \text{ mol}$$

Step 2: Find the ratio between magnesium and oxygen in the reaction

$$\text{Mg} : \text{O}_2 = 2 : 1$$

$$0.25 / 2 = 0.125$$

Step 3: Calculate the mass of oxygen that was reacted

$$\text{Mass} = \text{mol} \times \text{Mr}$$

$$\text{Mr of O}_2 = 32$$

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