

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and  $K_a$  or  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                        |                             |                   |
|------------------------|-----------------------------|-------------------|
| 1. autoionization      | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization          | 5. $K_a$                    | 6. $K_b$          |
| 7. ionization constant | 8. $pH$                     |                   |

## Key Equation

- |                       |                       |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

## Part A Completion

Write the equation and/or the numerical relationship of the values indicated that are indicated in the spaces provided. Write the equation and/or the numerical relationship.

1. When  $pH = 4$ , the concentration of  $H^+$  is \_\_\_\_\_.
2. The concentration of  $H^+$  is  $1 \times 10^{-4}$  M. The  $pH$  is \_\_\_\_\_.
3. The  $pH$  is 4. The concentration of  $H^+$  is \_\_\_\_\_.
4. The concentration of  $H^+$  is  $1 \times 10^{-4}$  M. The  $pH$  is \_\_\_\_\_.
5. The  $pH$  is 4. The concentration of  $H^+$  is \_\_\_\_\_.
6. The concentration of  $H^+$  is  $1 \times 10^{-4}$  M. The  $pH$  is \_\_\_\_\_.
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17. The  $pH$  is 4. The concentration of  $H^+$  is \_\_\_\_\_.
18. The concentration of  $H^+$  is  $1 \times 10^{-4}$  M. The  $pH$  is \_\_\_\_\_.
19. The  $pH$  is 4. The concentration of  $H^+$  is \_\_\_\_\_.
20. The concentration of  $H^+$  is  $1 \times 10^{-4}$  M. The  $pH$  is \_\_\_\_\_.

1.  $pH = 4$
2.  $[H^+] = 1 \times 10^{-4}$  M
3.  $pH = 4$
4.  $[H^+] = 1 \times 10^{-4}$  M
5.  $pH = 4$
6.  $[H^+] = 1 \times 10^{-4}$  M
7.  $pH = 4$
8.  $[H^+] = 1 \times 10^{-4}$  M
9.  $pH = 4$
10.  $[H^+] = 1 \times 10^{-4}$  M
11.  $pH = 4$
12.  $[H^+] = 1 \times 10^{-4}$  M
13.  $pH = 4$
14.  $[H^+] = 1 \times 10^{-4}$  M
15.  $pH = 4$
16.  $[H^+] = 1 \times 10^{-4}$  M
17.  $pH = 4$
18.  $[H^+] = 1 \times 10^{-4}$  M
19.  $pH = 4$
20.  $[H^+] = 1 \times 10^{-4}$  M

# 19 2 Hydrogen Ions And Acidity Answer Key

**Starr,Cecie Starr**



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