

AP World History Chapter 12 Latest 2022 Graded A+

Which of the following was a feature of the communities in Australia in the fifteenth century?

✓✓The practice of "fire stick farming"

Why did the maritime expeditions of the Indian Ocean basin sponsored by the Ming emperor suddenly stop in 1433? ✓✓The emperor's successors viewed expansion as a waste of resources.

In its efforts to recover from the disruption of Mongol rule, where did the Ming dynasty look for inspiration? ✓✓The culture of past Chinese dynasties

In which empire was a Shia version of Islam made the official religion in the sixteenth century?

✓✓Safavid Empire

Which of the following was a motivation for the European expansion but not for Chinese expansion in the fifteenth century? ✓✓To gain direct access to the goods and wealth of the Indian Ocean basin.

Which of the following is true of the Mughal Empire? ✓✓It established unified control over most of the Indian peninsula.

What feature did the Mughal Empire and Songhay Empire share? ✓✓The rulers were Muslim, but the majority of the population was not.

The wars carried out by the Aztec Empire were linked to... ✓✓the practice of Human sacrifice.

What did the Inca empire do that the Aztec empire did not do? ✓✓Build an elaborate bureaucracy to integrate and control its subjects.

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**Apex Learning, Mark Bach, Betsy
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incorrect because 566 kJ/mol is the Δh per mole of O₂ and not for 2 moles

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web thermochemical equations equations that represent both mass and enthalpy changes $h_2 + 2 O_2 \rightarrow 2 H_2O + 601 \text{ kJ/mol}$ this is an endothermic process it requires 601 kJ to melt one mole of ice $h_2 + 2 O_2 \rightarrow 2 H_2O + 601 \text{ kJ/mol}$ the enthalpy value will change if the number of moles varies from the 1:1 reaction stoichiometry this is an exothermic process

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web in this set of practice questions we will summarize the main concepts of thermochemistry such as the relationship between internal energy work and heat exothermic and endothermic process heat capacity constant pressure calorimetry constant volume calorimetry the enthalpy the standard enthalpies of formation and their use in

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy ΔE of a system in any process is equal to the heat Q added to the system plus the work W done on the system by its surroundings $\Delta E = Q + W$

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