

## AP CHEMISTRY MULTIPLE CHOICE QUESTIONS

## Unit 8 – Introduction to Equilibrium



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



Reversed  
+ Doubled

(A) 95<sup>2</sup>

(B) 95

(C) (95)<sup>1/2</sup>

(D)  $\frac{1}{95}$

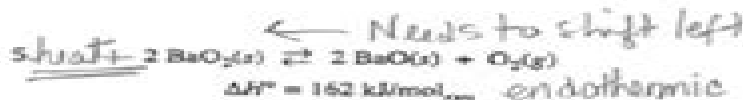
(E)  $\frac{1}{(95)^2}$



$C_2H_4(g)$  is reduced by  $H_2(g)$  in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of  $H_2(g)$
- II. Changes in the particle size of the platinum catalyst
- III. Changes in the temperature of the reaction system

- (A) III only
- (B) I and II only
- (C) I and III only
- (D) II and III only
- (E) I, II, and III



A sealed rigid vessel contains  $BaO_2(s)$  in equilibrium with  $BaO(s)$  and  $O_2(g)$  as represented by the equation above. Which of the following changes will increase the amount of  $BaO_2(s)$  in the vessel?

- (A) Removing a small amount of  $O_2(g)$   $R \times$
- (B) Removing a small amount of  $BaO(s)$   $X$
- (C) Adding He gas to the vessel  $X$
- (D) Lowering the temperature  $L$

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products?

increase  $P^\circ$

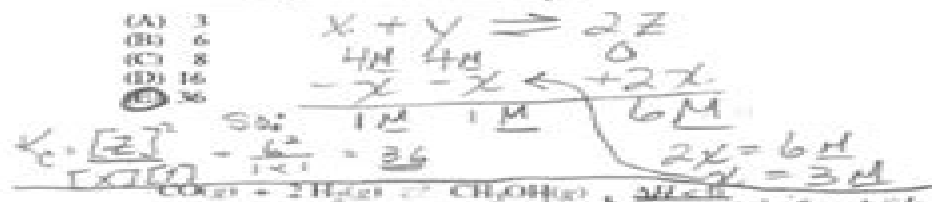


shift towards fewer  
+ make gas



When 4.00 mol each of  $X(g)$  and  $Y(g)$  are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of  $Z(g)$  is produced. What is the value of the equilibrium constant,  $K_c$ ?

- (A) 3
- (B) 6
- (C) 8
- (D) 16
- (E) 36



6. The synthesis of  $CH_3OH(l)$  from  $CO_2(g)$  and  $H_2(g)$  is represented by the equation above. The value of  $K_c$  for the reaction at 483 K is 14.5.

increase heat shifts left, decrease  $K_c$   
it is negative

Which of the following explains the effect on the equilibrium constant,  $K_c$ , when the temperature of the reaction system is increased to 650 K?

- (A)  $K_c$  will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.
- (B)  $K_c$  will increase because there are more reactant molecules than product molecules.
- (C)  $K_c$  will decrease because the reaction is exothermic.
- (D)  $K_c$  is constant and will not change.

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**Linda D. Williams**



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