

**AP CHEMISTRY MULTIPLE CHOICE QUESTIONS**  
**Unit 8 – Introduction to Equilibrium**



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



$$\begin{array}{l} \text{(A) } 0.95 \text{ g/cm}^2 \\ \text{(B) } 9.5 \text{ g/cm}^2 \\ \text{(C) } 0.95 \text{ g/cm}^3 \\ \text{(D) } \frac{1}{0.95} \text{ g/cm}^2 \end{array}$$



$C_2H_6O_2$ ( $CH_3COOH$ ) is reduced by  $H_2(g)$  in the presence of a solid platinum catalyst, as represented by the equation above. Factors that would affect the rate of the reaction include which of the following?

I. Changes in the partial pressure of  $\text{H}_2\text{O(g)}$   
 II. Changes in the particle size of the platinum catalyst  
 III. Changes in the temperature of the reaction system

(A) III only  
 (B) I and II only  
 (C) I and III only  
 (D) II and III only  
 (E) I, II, and III

$$\text{shift} \leftarrow \text{N} \rightarrow \text{shift left}$$

$$\text{shift} + 2 \text{BaO}_2(s) \rightleftharpoons 2 \text{BaO}(s) + \text{O}_2(g)$$



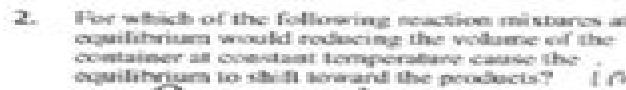
A sealed rigid vessel contains  $\text{BaO}_2(s)$  in equilibrium with  $\text{BaO}(s)$  and  $\text{O}_2(g)$  as represented by the equation above. Which of the following changes will increase the amount of  $\text{BaO}_2(s)$  in the vessel?

(A) Removing a small amount of  $O_2(g)$

(B) Removing a small amount of  $Ba(OH)_2$

(C) Adding  $H_2$  gas to the vessel

(D) Lowering the temperature



$$\begin{array}{l}
 \text{(A) } \text{SrCO}_3(s) \rightleftharpoons \text{SrO}(s) + \text{CO}_2(g) \\
 \text{(B) } \text{CO}(g) + \text{O}_2(g) \rightleftharpoons 2\text{CO}_2(g) \\
 \text{(C) } \text{CO}_2(g) + \text{H}_2\text{O}(l) \rightleftharpoons \text{H}_2\text{CO}_3(aq) \\
 \text{(D) } \text{PCl}_5(s) \rightleftharpoons \text{PCl}_3(g) + \text{Cl}_2(g)
 \end{array}$$

(E) 2 factors = natural shift towards it makes fewer gas



When 4.00 mol each of  $X(g)$  and  $Y(g)$  are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of  $Z(g)$  is produced. What is the value of the equilibrium constant,  $K$ ?

(A) 3  
 (B) 6  
 (C) 8  
 (D) 16  
 (E) 26

$$\begin{array}{r}
 X + Y = 22 \\
 4M \quad 4M \quad 0 \\
 -X \quad -X \leftarrow +2X \\
 \hline
 1M \quad 1M \quad 6M
 \end{array}$$

6. The synthesis of  $\text{CH}_3\text{OH}(g)$  from  $\text{CO}(g)$  and  $\text{H}_2(g)$  is represented by the equation above. The value of  $K_c$  is negligible.



Which of the following explains the effect on the equilibrium constant,  $K_c$ , when the temperature of the reaction system is increased to 650 K?

(A)  $K_c$  will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.

(B)  $K_c$  will increase because there are more reactant molecules than product molecules.

(C)  $K_c$  will decrease because the reaction is exothermic.

(D)  $K_c$  is constant and will not change.

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**Linda D. Williams**

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