

Matter—Properties and Changes

Section 3.1 Properties of Matter

pages 70–75

Problem-Solving Lab

- Explain why the flow of a compressed gas must be controlled for practical and safe use.
The flow of compressed gas must be controlled to control the amount and the rate at which gas is released.
- Predict what would happen if the valve on a full tank of compressed gas were suddenly opened all the way or if the tank were accidentally punctured.

Without the regulator device, the gas would rush out of the tank with a force powerful enough to transform the tank into a dangerous, uncontrolled projectile.

Section 3.1 Assessment

page 75

- Create a table that describes the three common states of matter in terms of their shape, volume, and compressibility.

	Volume	Shape	Compressibility
Solid	Definite	Definite	Incompressible
Liquid	Definite	Takes shape of container and fills container to the extent of its own volume	Virtually incompressible
Gas	Fills volume of container	Takes shape of container	Compressible

- Describe the characteristics that identify a sample of matter as being a substance.
The sample of matter must have a uniform and unchanging composition to be a substance.
- Classify each of the following as a physical or chemical property.
 - Iron and oxygen form rust.
chemical

- Iron is more dense than aluminum.
physical
- Magnesium burns brightly when ignited.
chemical
- Oil and water do not mix.
physical
- Mercury melts at -39°C .
physical

- Organize Create a chart that compares physical and chemical properties. Give two examples for each type of property.

The chart should make clear that physical properties can be observed without changing the composition of the sample, which is not the case for chemical properties. Mass and density are examples of physical properties. Fermentation and rusting are examples of chemical properties.

Section 3.2 Changes in Matter

pages 76–79

Practice Problems

page 78

- Use the data in the table to answer the following questions.

Aluminum and Liquid Bromine Reaction

	Before Reaction	After Reaction
Aluminum	10.3 g	0.0 g
Liquid bromine	100.0 g	8.5 g
Compound	0.0 g	

How many grams of bromine reacted? How many grams of compound were formed?

amount of bromine that reacted =
 $100.0 \text{ g} - 8.5 \text{ g} = 91.5 \text{ g}$

amount of compound formed =
 $100.0 \text{ g} + 10.3 \text{ g} - 8.5 \text{ g} = 101.8 \text{ g}$

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Table of Contents Section 3 1 Chemistry Answers

1. Understanding the eBook Section 3 1 Chemistry Answers
 - The Rise of Digital Reading Section 3 1 Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 3 1 Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 3 1 Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 3 1 Chemistry Answers
 - Personalized Recommendations
 - Section 3 1 Chemistry Answers User Reviews and Ratings
 - Section 3 1 Chemistry Answers and Bestseller Lists
5. Accessing Section 3 1 Chemistry Answers Free and Paid eBooks
 - Section 3 1 Chemistry Answers Public Domain eBooks
 - Section 3 1 Chemistry Answers eBook Subscription Services
 - Section 3 1 Chemistry Answers Budget-Friendly Options
6. Navigating Section 3 1 Chemistry Answers eBook Formats

- ePub, PDF, MOBI, and More
- Section 3 1 Chemistry Answers Compatibility with Devices
- Section 3 1 Chemistry Answers Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Section 3 1 Chemistry Answers
- Highlighting and Note-Taking Section 3 1 Chemistry Answers
- Interactive Elements Section 3 1 Chemistry Answers

8. Staying Engaged with Section 3 1 Chemistry Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Section 3 1 Chemistry Answers

9. Balancing eBooks and Physical Books Section 3 1 Chemistry Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Section 3 1 Chemistry Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Section 3 1 Chemistry Answers

- Setting Reading Goals Section 3 1 Chemistry Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Section 3 1 Chemistry Answers

- Fact-Checking eBook Content of Section 3 1 Chemistry Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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