

Acids and Bases Worksheet

1. Acids and bases share some properties but are different in others. Use the samples at the side of the class to complete the chart. If both acids and bases share the property, place it in both columns.

Property	Acids	Bases
Conductivity in Aqueous Solution	Very	Very
Taste	Sour	Bitter
Feel	—	Slippery
Reaction with Metal	produces H_2 gas	—
Reaction with Carbonate	produces CO_2 gas	—
Colour on Litmus Paper	Red	Blue
Colour with Universal Indicator	Red	Purple
Colour with phenolphthalein	Colourless	Pink

2. Indicate if each compound is an acid or base and name it.

	Acid/Base?	Name
HNO_3	A	nitric acid
$Mg(OH)_2$	B	magnesium hydroxide
H_3PO_4	A	phosphoric acid
H_2CO_3	A	carbonic acid
HF	A	hydrofluoric acid
$Fe(OH)_3$	B	iron (III) hydroxide

3. Write the formula for each of the following acids/bases:

hydrochloric acid	<u>HCl</u>	hydrochloric acid	<u>HCl</u>
aluminum hydroxide	<u>$Al(OH)_3$</u>	sodium hydroxide	<u>$NaOH$</u>
lithium hydroxide	<u>$LiOH$</u>	calcium hydroxide	<u>$Ca(OH)_2$</u>
sulfuric acid	<u>H_2SO_4</u>	sulfurous acid	<u>H_2SO_3</u>

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J Elliott

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The Chemistry of the Several Natural and Artificial Heterogeneous Compounds Used in Manufacturing Porcelain, Glass, and Pottery Simeon Shaw, 1900 **Prentice Hall Chemistry**, 2000 **Chemistry Textbook for College and University**

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Table of Contents 19 3 Acids And Bases Answer Key Chemistry

1. Understanding the eBook 19 3 Acids And Bases Answer Key Chemistry
 - The Rise of Digital Reading 19 3 Acids And Bases Answer Key Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Acids And Bases Answer Key Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 3 Acids And Bases Answer Key Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Acids And Bases Answer Key Chemistry
 - Personalized Recommendations
 - 19 3 Acids And Bases Answer Key Chemistry User Reviews and Ratings
 - 19 3 Acids And Bases Answer Key Chemistry and Bestseller Lists
5. Accessing 19 3 Acids And Bases Answer Key Chemistry Free and Paid eBooks
 - 19 3 Acids And Bases Answer Key Chemistry Public Domain eBooks
 - 19 3 Acids And Bases Answer Key Chemistry eBook Subscription Services
 - 19 3 Acids And Bases Answer Key Chemistry Budget-Friendly Options

6. Navigating 19 3 Acids And Bases Answer Key Chemistry eBook Formats
 - ePUB, PDF, MOBI, and More
 - 19 3 Acids And Bases Answer Key Chemistry Compatibility with Devices
 - 19 3 Acids And Bases Answer Key Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 3 Acids And Bases Answer Key Chemistry
 - Highlighting and Note-Taking 19 3 Acids And Bases Answer Key Chemistry
 - Interactive Elements 19 3 Acids And Bases Answer Key Chemistry
8. Staying Engaged with 19 3 Acids And Bases Answer Key Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 3 Acids And Bases Answer Key Chemistry
9. Balancing eBooks and Physical Books 19 3 Acids And Bases Answer Key Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 3 Acids And Bases Answer Key Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 19 3 Acids And Bases Answer Key Chemistry
 - Setting Reading Goals 19 3 Acids And Bases Answer Key Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 19 3 Acids And Bases Answer Key Chemistry
 - Fact-Checking eBook Content of 19 3 Acids And Bases Answer Key Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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