

# Chemistry Study Guide

### Periodic Table of the Elements

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<div><div>Li</div><div>3</div><div>6.941</div></div>	<div><div>Be</div><div>4</div><div>9.012</div></div>																	<div><div>B</div><div>5</div><div>10.811</div></div>	<div><div>C</div><div>6</div><div>12.011</div></div>	<div><div>N</div><div>7</div><div>14.007</div></div>	<div><div>O</div><div>8</div><div>15.999</div></div>	<div><div>F</div><div>9</div><div>18.998</div></div>	<div><div>Ne</div><div>10</div><div>20.180</div></div>
<div><div>Na</div><div>11</div><div>22.990</div></div>	<div><div>Mg</div><div>12</div><div>24.305</div></div>																	<div><div>Al</div><div>13</div><div>26.982</div></div>	<div><div>Si</div><div>14</div><div>28.086</div></div>	<div><div>P</div><div>15</div><div>30.974</div></div>	<div><div>S</div><div>16</div><div>32.065</div></div>	<div><div>Cl</div><div>17</div><div>35.453</div></div>	<div><div>Ar</div><div>18</div><div>39.948</div></div>
<div><div>K</div><div>19</div><div>39.098</div></div>	<div><div>Ca</div><div>20</div><div>40.078</div></div>	<div><div>Sc</div><div>21</div><div>44.956</div></div>	<div><div>Ti</div><div>22</div><div>47.88</div></div>	<div><div>V</div><div>23</div><div>50.942</div></div>	<div><div>Cr</div><div>24</div><div>51.996</div></div>	<div><div>Mn</div><div>25</div><div>54.938</div></div>	<div><div>Fe</div><div>26</div><div>55.845</div></div>	<div><div>Co</div><div>27</div><div>58.933</div></div>	<div><div>Ni</div><div>28</div><div>58.693</div></div>	<div><div>Cu</div><div>29</div><div>63.546</div></div>	<div><div>Zn</div><div>30</div><div>65.38</div></div>	<div><div>Ga</div><div>31</div><div>69.723</div></div>	<div><div>Ge</div><div>32</div><div>72.64</div></div>	<div><div>As</div><div>33</div><div>74.922</div></div>	<div><div>Se</div><div>34</div><div>78.96</div></div>	<div><div>Br</div><div>35</div><div>79.904</div></div>	<div><div>Kr</div><div>36</div><div>83.80</div></div>						
<div><div>Rb</div><div>37</div><div>85.468</div></div>	<div><div>Sr</div><div>38</div><div>87.62</div></div>	<div><div>Y</div><div>39</div><div>88.906</div></div>	<div><div>Zr</div><div>40</div><div>91.224</div></div>	<div><div>Nb</div><div>41</div><div>92.906</div></div>	<div><div>Mo</div><div>42</div><div>95.94</div></div>	<div><div>Tc</div><div>43</div><div>98.00</div></div>	<div><div>Ru</div><div>44</div><div>101.07</div></div>	<div><div>Rh</div><div>45</div><div>102.91</div></div>	<div><div>Pd</div><div>46</div><div>106.42</div></div>	<div><div>Ag</div><div>47</div><div>107.87</div></div>	<div><div>Cd</div><div>48</div><div>112.41</div></div>	<div><div>In</div><div>49</div><div>114.82</div></div>	<div><div>Sn</div><div>50</div><div>118.71</div></div>	<div><div>Sb</div><div>51</div><div>121.76</div></div>	<div><div>Te</div><div>52</div><div>127.6</div></div>	<div><div>I</div><div>53</div><div>126.91</div></div>	<div><div>Xe</div><div>54</div><div>131.29</div></div>						
<div><div>Cs</div><div>55</div><div>132.91</div></div>	<div><div>Ba</div><div>56</div><div>137.33</div></div>																	<div><div>Tl</div><div>81</div><div>204.38</div></div>	<div><div>Pb</div><div>82</div><div>207.2</div></div>	<div><div>Bi</div><div>83</div><div>208.98</div></div>	<div><div>Po</div><div>84</div><div>209</div></div>	<div><div>At</div><div>85</div><div>210</div></div>	<div><div>Rn</div><div>86</div><div>222</div></div>
<div><div>Fr</div><div>87</div><div>223</div></div>	<div><div>Ra</div><div>88</div><div>226</div></div>																	<div><div>Uut</div><div>113</div><div>288</div></div>	<div><div>Fl</div><div>114</div><div>289</div></div>	<div><div>Uup</div><div>115</div><div>288</div></div>	<div><div>Lv</div><div>116</div><div>293</div></div>	<div><div>Uus</div><div>117</div><div>294</div></div>	<div><div>Uuo</div><div>118</div><div>294</div></div>
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## Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

## Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

## Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	$p^+$	1	+1	Nucleus
Neutron	$n$	1	0	Nucleus
Electron	$e^-$	$\sim 0$	-1	Orbit/Shell

# Section 2 nuclear Chemistry Study Guide Key

**SILBERBERG**



## Section 2nuclear Chemistry Study Guide Key:

**Study Guide with Answers to Selected Problems** H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises Chemistry, Student Study Guide John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that s real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You ll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors **Study Guide, Answer Book to Accompany Organic Chemistry** Alan S. Wingrove, Robert L. Caret, 1981 *General Chemistry Key Concise Review Guide Sixth Edition* Darrell D. Ebbing, 1999 Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter **Study Guide and Partial Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry** Susan McMurry, 1996 This internationally acclaimed detective series is just the thing for lovers of those Number One Ladies looking for a darker more realistic view of Botswana Sue Baker Publishing News **Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition** Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Price winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every

three chapters more detailed explanations in solutions and chapter outlines *Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e* Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry **General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999 **Study Guide with Solutions for Organic Chemistry, Third Edition** Ralph J. Fessenden, Joan S. Fessenden, 1986 Provides a set of additional drill problems chapter by chapter discussions and supplemental instructional material to help students master organic chemistry problem solving techniques **Chemistry** Charles E. Mortimer, 1983 Good No Highlights No Markup all pages are intact Slight Shelfwear may have the corners slightly dented may have slight color changes slightly damaged spine

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