

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain |
| 7. autoionization | 8. acid rain | 9. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation for the autoionization of water. What is the K_w for water at 25°C? What is the K_w for water at 100°C? What is the K_w for water at 0°C?

1. What is the K_w for water at 25°C? What is the K_w for water at 100°C? What is the K_w for water at 0°C?
2. What is the K_w for water at 25°C? What is the K_w for water at 100°C? What is the K_w for water at 0°C?
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9. What is the K_w for water at 25°C? What is the K_w for water at 100°C? What is the K_w for water at 0°C?
10. What is the K_w for water at 25°C? What is the K_w for water at 100°C? What is the K_w for water at 0°C?

1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2. 1.0×10^{-14}
3. 1.0×10^{-14}
4. 1.0×10^{-14}
5. 1.0×10^{-14}
6. 1.0×10^{-14}
7. 1.0×10^{-14}
8. 1.0×10^{-14}
9. 1.0×10^{-14}
10. 1.0×10^{-14}

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Albert A Gayle



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