

Chemistry Study Guide

Periodic Table of the Elements

<small>1</small> H <small>1.008</small>																	<small>2</small> He <small>4.003</small>																															
<small>3</small> Li <small>6.941</small>	<small>4</small> Be <small>9.012</small>																	<small>5</small> B <small>10.81</small>	<small>6</small> C <small>12.01</small>	<small>7</small> N <small>14.01</small>	<small>8</small> O <small>16.00</small>	<small>9</small> F <small>18.998</small>	<small>10</small> Ne <small>20.18</small>																									
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<small>19</small> K <small>39.10</small>	<small>20</small> Ca <small>40.08</small>	<small>21</small> Sc <small>44.96</small>	<small>22</small> Ti <small>47.88</small>	<small>23</small> V <small>50.94</small>	<small>24</small> Cr <small>52.00</small>	<small>25</small> Mn <small>54.94</small>	<small>26</small> Fe <small>55.85</small>	<small>27</small> Co <small>58.93</small>	<small>28</small> Ni <small>58.71</small>	<small>29</small> Cu <small>63.55</small>	<small>30</small> Zn <small>65.38</small>	<small>31</small> Ga <small>69.72</small>	<small>32</small> Ge <small>72.64</small>	<small>33</small> As <small>74.92</small>	<small>34</small> Se <small>78.96</small>	<small>35</small> Br <small>79.90</small>	<small>36</small> Kr <small>83.80</small>																															
<small>37</small> Rb <small>85.47</small>	<small>38</small> Sr <small>87.62</small>	<small>39</small> Y <small>88.91</small>	<small>40</small> Zr <small>91.22</small>	<small>41</small> Nb <small>92.91</small>	<small>42</small> Mo <small>95.94</small>	<small>43</small> Tc <small>98.91</small>	<small>44</small> Ru <small>101.07</small>	<small>45</small> Rh <small>102.91</small>	<small>46</small> Pd <small>106.42</small>	<small>47</small> Ag <small>107.87</small>	<small>48</small> Cd <small>112.41</small>	<small>49</small> In <small>114.82</small>	<small>50</small> Sn <small>118.71</small>	<small>51</small> Sb <small>121.76</small>	<small>52</small> Te <small>127.60</small>	<small>53</small> I <small>126.91</small>	<small>54</small> Xe <small>131.29</small>																															
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Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

Section 2 nuclear Chemistry Study Guide Key

John McMurry, Susan McMurry



Section 2nuclear Chemistry Study Guide Key:

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- Chapter 2: Essential Elements of Section 2nuclear Chemistry Study Guide Key
- Chapter 3: Section 2nuclear Chemistry Study Guide Key in Everyday Life
- Chapter 4: Section 2nuclear Chemistry Study Guide Key in Specific Contexts
- Chapter 5: Conclusion

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3. In chapter 2, the author will delve into the foundational concepts of Section 2nuclear Chemistry Study Guide Key. The second chapter will elucidate the essential principles that must be understood to grasp Section 2nuclear Chemistry Study Guide Key in its entirety.
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