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John Walkenbach



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 copper which is put in the calorimeter are the same the initial temperature of the calorimeter and water is 30 c and that of
 hot object is 60 c the specific heat of copper and water are 0 09 cal gm c and 1 cal gm c respectively

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 temperature change plugging in the values given in the problem you get $q_{\text{water}} = 4.18 \text{ J/g} \cdot \text{C} \times 110 \text{ g} \times 26.6 \text{ C} = 1250 \text{ J}$
 $550 \text{ J} = \Delta h_{\text{q water}} = 550 \text{ J}$

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 for water 4.184 J/g°C 1 what is the specific heat of aluminum if the temperature of a 28.4 g sample of aluminum is increased
 by 8.1°C when 207 J of heat is added 2

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 $14.73 \text{ kJ} = \Delta H_{\text{rxn}} = q_{\text{rxn}} = q_{\text{calorimeter}} = mcs\Delta t$

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