

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from pH and pOH , and vice versa from pH and pOH .
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. pH | 6. pOH |

Key Equation

- | | |
|-------------------------------------|---|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = \text{pH} + \text{pOH}$ |

Part A Completion

Write the requested unit or numerical information of the values provided that are indicated in parentheses. Significant figures are indicated with the appropriate symbol.

1. What substance is _____, a weak hydrogen ion concentration of _____? The concentration of this acidic substance is _____.
2. Hydrochloric acid has a pH of _____ or a _____ molar concentration. _____ concentration is _____ the _____ strength of _____.
3. _____ is a _____ acid. _____ is a _____ acid. _____ is a _____ acid.
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Wilhelm Mansfield Clark



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