

# **ACS ANALYTICAL CHEMISTRY STUDY GUIDE**

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- $k_f'$ :  $k_f' = k_f^* \cdot a$
- EDTA titrations: 1. before eq EDTA binds to metal 1:1
- 2. At eq we assume that all metal ions can bind to EDTA and then use  $k_f'$
- Potentiometry: an analytical method in which an electric potential difference (avoltage) of a cell is measured
  
- Reducing agent: Donates electrons and becomes oxidized.
- oxidizing agent: Accepts electrons and becomes reduced.
- coulometry: science of measuring change and relating it to analyte
- cathode: the electrode at which reduction occurs
- Anode: the electrode at which oxidation occurs
- Ecell:  $E_{cell} = E_{red} - E_{ox}$
- Faraday's Equation:  $E = E - (0.0592/n) \log(a/b)$
- ||: anode always written on left of ||
- amperometry: elec techniques where the cell current is measured where pot diff between the indicator and reference electrode is controlled
- Clark electrode: Electrode for oxygen
- Biosensors: Detecting devices that can detect specific substances in cells
- amp chromo det: point, size, or charge
- limiting current: the diff between current and steady state max
- Anodic vol: Very sensative for analyzing metals into ppb
- fluroescence: used to measure singlets, atoms like rigid aromatic rings, high sensitivity
- phosphorescence: involves emission of light trip- singlet, much lower than fluro
- Transmittance:  $A = -\log(T)$  or  $2 - \log(T\%)$
- atomic absorbtion: lamp-flame-mono-detector
- atomic emission: flame-mono-detector
- A.A. hollow cathode lamp: used to form very narrow band lights
- A.A. atomizer: Heat sample to very high temps to break all chemical bonds, hot enough to

# Acs National Chemistry Exam Study Guide

**Barbara L. Yoost, Lynne R.  
Crawford, Patricia Castaldi**

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