

## Acids, Bases and Salts

- THE THEORY of ACIDS and ALKALIS and a few technical terms:
  - Acids are substances that form hydrogen ions ( $H^{+}$ ) when dissolved in water e.g. hydrochloric acid HCl gives  $H^{+}$  and  $Cl^{-}$  ions, sulphuric acid  $H_2SO_4$  gives  $2H^{+}$  and  $SO_4^{2-}$  ions and nitric acid  $HNO_3$  gives  $H^{+}$  and  $NO_3^{-}$  ions.
  - Alkalies are substances that form hydroxide ions ( $OH^{-}$ ) in water e.g. sodium hydroxide NaOH gives  $Na^{+}$  and  $OH^{-}$  ions, calcium hydroxide  $Ca(OH)_2$  gives  $Ca^{2+}$  and  $2OH^{-}$  ions. Note: an alkali is a base soluble in water.
  - In water, there are trace quantities of  $H^{+}$  and  $OH^{-}$  ions BUT they are of equal concentration and so water is neutral.
  - In acid solutions there are more  $H^{+}$  ions than  $OH^{-}$  ions.
  - In alkaline solution there are more  $OH^{-}$  ions than  $H^{+}$  ions.
  - Acids dissociate to different extents in aqueous solution. Acids that dissociate to a large extent are strong electrolytes and strong acids. In contrast, acids that dissociate only to a small extent are weak acids and weak electrolytes.

In a similar manner, bases can be strong or weak depending on the extent to which they dissociate and produce  $OH^{-}$  ions in solution. Most metal hydroxides are strong electrolytes and strong bases. Ammonia,  $NH_3$ , is a weak electrolyte and weak base.



- BASES e.g. oxides and hydroxides are substances that react and neutralise acids to form salts and water. Bases which are soluble in water are called alkalis.

### Acids

Some common acids are listed below:

Name	Formula	Strong/Weak	Where is it found?
Hydrochloric acid	HCl	Strong	The stomach, in the lab.
Sulphuric acid	$H_2SO_4$	Strong	Acid rain, car batteries, the lab.
Nitric acid	$HNO_3$	Strong	Acid rain, in the lab.
Ethanoic (acetic) acid	$CH_3COOH$	Weak	Vinegar
Methanoic (formic) acid	$HCOOH$	Weak	Ant & nettle stings, descalers
Citric Acid	$C_6H_8O_7$	Weak	Citrus fruits

- Acids taste sour (e.g. vinegar, lemon juice).
- Acids are harmful to living cells.
- Aqueous solutions of all acids contain hydrogen ions,  $H^{+}$ .

# Acids Bases And Salts Study Guide

**Steven S. Zumdahl, Paul B. Kelter**

## Acids Bases And Salts Study Guide:

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Henold,1984    **A Study Guide to Chemical Principles** Wilbert Hutton,1970    *Study Guide for General Chemistry for Colleges* Herman Thompson Briscoe,1936    Fundamentals of Medicinal Chemistry and Drug Metabolism M. O. Faruk Khan,Ashok Philip,2018-06-01 The primary objective of this 4 volume book series is to educate PharmD students on the subject of medicinal chemistry The book set serves as a reference guide to pharmacists on aspects of chemical basis of drug action This first volume of the series is comprised of 8 chapters focusing on basic background information about medicinal chemistry It takes a succinct and conceptual approach to introducing important fundamental concepts required for a clear understanding of various facets of pharmacotherapeutic agents drug metabolism and important biosynthetic pathways that are relevant to drug action Notable topics covered in this first volume include the scope and importance of medicinal chemistry in pharmacy education a comprehensive discussion of the organic functional groups present in drugs and information about four major types of biomolecules proteins carbohydrates lipids nucleic acids and key heterocyclic ring systems The concepts of acid base chemistry and salt formation and their applications to the drug action and design follow thereafter These include concepts of solubility and lipid water partition coefficient LWPC isosterism stereochemical properties mechanisms of drug action drug receptor interactions critical for pharmacological responses of drugs and much more Students and teachers will be able to integrate the knowledge presented in the book and apply medicinal chemistry concepts to understand the pharmacodynamics and pharmacokinetics of therapeutic agents in the body

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## **Table of Contents Acids Bases And Salts Study Guide**

1. Understanding the eBook Acids Bases And Salts Study Guide
  - The Rise of Digital Reading Acids Bases And Salts Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Acids Bases And Salts Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acids Bases And Salts Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Acids Bases And Salts Study Guide
  - Personalized Recommendations
  - Acids Bases And Salts Study Guide User Reviews and Ratings
  - Acids Bases And Salts Study Guide and Bestseller Lists
5. Accessing Acids Bases And Salts Study Guide Free and Paid eBooks
  - Acids Bases And Salts Study Guide Public Domain eBooks
  - Acids Bases And Salts Study Guide eBook Subscription Services
  - Acids Bases And Salts Study Guide Budget-Friendly Options
6. Navigating Acids Bases And Salts Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Acids Bases And Salts Study Guide Compatibility with Devices
  - Acids Bases And Salts Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Acids Bases And Salts Study Guide
  - Highlighting and Note-Taking Acids Bases And Salts Study Guide
  - Interactive Elements Acids Bases And Salts Study Guide
8. Staying Engaged with Acids Bases And Salts Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Acids Bases And Salts Study Guide

9. Balancing eBooks and Physical Books Acids Bases And Salts Study Guide

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Acids Bases And Salts Study Guide

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Acids Bases And Salts Study Guide

- Setting Reading Goals Acids Bases And Salts Study Guide
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Acids Bases And Salts Study Guide

- Fact-Checking eBook Content of Acids Bases And Salts Study Guide
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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